

**Assignment #1 – Compound Names and Formulas**  
**Single-valent ions only**

**A. Name these Compounds**

1.  $\text{Li}_2\text{S}$  Lithium Sulfide
2.  $\text{CaO}$  Calcium Oxide
3.  $\text{NaF}$  Sodium Fluoride
4.  $\text{CaBr}_2$  Calcium Bromide
5.  $\text{MgCl}_2$  Magnesium Chloride
6.  $\text{BBr}_3$  Boron Bromide
7.  $\text{Cs}_2\text{O}$  Cesium Oxide
8.  $\text{FrBr}$  Francium Bromide
9.  $\text{Ag}_2\text{S}$  Silver Sulfide

10.  $\text{GeF}_4$  Germanium Fluoride
11.  $\text{Ga}_2\text{O}_3$  Gallium Oxide
12.  $\text{EsCl}_3$  Einsteinium Chloride
13.  $\text{Fm}_2\text{O}_3$  Fermium Oxide
14.  $\text{Mg}_3\text{N}_2$  Magnesium Nitride
15.  $\text{Rb}_2\text{O}$  Rubidium Oxide
16.  $\text{RaO}$  Radium Oxide
17.  $\text{SrO}$  Strontium Oxide
18.  $\text{Tc}_2\text{O}_7$  Technetium Oxide

**B. Write the correct chemical formula for these compounds by balancing the ionic charges**

1. sodium chloride  $\text{NaCl}$
2. magnesium fluoride  $\text{MgF}_2$
3. silver oxide  $\text{Ag}_2\text{O}$
4. indium bromide  $\text{InBr}_3$
5. zinc bromide  $\text{ZnBr}_2$
6. neodymium oxide  $\text{Nd}_2\text{O}_3$
7. thorium sulphide  $\text{ThS}_2$
8. actinium oxide  $\text{Ac}_2\text{O}_3$
9. radium bromide  $\text{RaBr}_2$
10. cesium oxide  $\text{Cs}_2\text{O}$

11. hydrogen oxide  $\text{H}_2\text{O}$
12. francium nitride  $\text{Fr}_3\text{N}$
13. rubidium phosphide  $\text{Rb}_3\text{P}$
14. potassium oxide  $\text{K}_2\text{O}$
15. beryllium sulphide  $\text{BeS}$
16. lithium sulphide  $\text{Li}_2\text{S}$
17. hydrogen bromide  $\text{HBr}$
18. strontium nitride  $\text{Sr}_3\text{N}_2$
19. calcium oxide  $\text{CaO}$
20. tantalum nitride  $\text{Ta}_3\text{N}_5$

# Writing And Naming Binary Compounds Worksheet

## Answer Key

**TD Snyder**



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