

Experiment 8-18

First Law of Thermodynamics: The Mechanical Equivalent of Heat

I. Synopsis

Students explore the equivalence of mechanical and heat energy through a series of experiments designed to determine the mechanical equivalent of heat.

II. Objectives

- Understand the concept of mechanical work and its relation to energy.
- Understand the concept of heat and its relation to energy.
- Understand the concept of the mechanical equivalent of heat.

III. Prerequisite

- Students should have completed the previous experiments in the course.
- Students should have completed the previous experiments in the course.

IV. Preliminary Questions

What is the relationship between the mechanical work done on a system and the heat added to the system? How is this relationship related to the First Law of Thermodynamics?

$$W = \int \mathbf{F} \cdot d\mathbf{r}$$

$$Q = \int \mathbf{F} \cdot d\mathbf{r}$$

where W is the work done on the system, Q is the heat added to the system, $\mathbf{F}$ is the force applied to the system, and $d\mathbf{r}$ is the displacement of the system. The relationship between W and Q is the mechanical equivalent of heat, which is the ratio of the mechanical work done on the system to the heat added to the system. This ratio is constant for all systems and is denoted by the symbol J .

First Law Of Thermodynamics Lab Report

DJ Losen



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