

- b.  $[\text{M}(\text{H}_2\text{O})_6]^{2+}$  **M** is second row transition metal, LFSE =  $-2.4 \Delta_o$ . This can be achieved with a low spin  $d^6$  configuration.  $\text{M}^{2+}$  with 6  $d$  electrons: **M** = Rh



- c.  $[\text{MCl}_4]^-$  **M** is first row transition metal, 5 unpaired electrons.  $\text{M}^{3+}$  with 5  $d$  electrons: Fe



- d.  $\text{MCl}_2(\text{NH}_3)_2$  **M** is third row  $d^8$  transition metal; two **M**–Cl stretching bands in IR. Second and third row  $d^8$  complexes are often square planar. The presence of two **M**–Cl stretching bands implies *cis* geometry of the chloro ligands.

$\text{M}^{2+}$  with 8  $d$  electrons: **M** = Pt



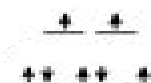
- 10.4** Angular overlap calculations for  $d^8$  and  $d^9$  ions show no energy difference between  $D_{4h}$  and  $O_h$  when exclusively considering  $\sigma$  interactions. Both  $d^8$  geometries have energies of  $-3e_{\sigma}$ ; both  $d^9$  geometries have energies of  $-6e_{\sigma}$ . In general, stability constants decrease as more ligands are added; the sequence for nickel is the common one. The huge drop in stability constant between the second and third ethylenediamine on  $\text{Cu}^{2+}$  is a result of the  $d^9$  Jahn-Teller effect. The first two en ligands add in a square-planar geometry, with water molecules in the axial positions, and the coordination of these two monodentate ligands allows for the Jahn-Teller distortion. Adding a third en ligand requires a geometry change and with a preference for uniform **M**–N bond distances towards the six nitrogen atoms. This is counter to the Jahn-Teller distortion, and the third addition is much less favorable than the first two.

- 10.5**  $[\text{M}(\text{H}_2\text{O})_6]^{2+}$  **M** is first row transition metal,  $\mu = 3.9 \text{ BM}$ . The magnetic moment implies 3 unpaired electrons, which would give rise to  $\mu = \sqrt{3(3+5)} = 3.9 \text{ BM}$ . There are two possibilities,  $d^3$  and  $d^7$ :

$\text{M}^{2+}$ ,  $d^3$ : **M** = V



$\text{M}^{2+}$ ,  $d^7$ : **M** = Co



- 10.6** a.  $[\text{Cr}(\text{H}_2\text{O})_6]^{2+}$   $n = 4$   
 $\mu = \sqrt{4(6)} = 4.9 \mu_B$

- b.  $[\text{Cr}(\text{CN})_6]^{4-}$   $n = 2$   
 $\mu = \sqrt{2(4)} = 2.8 \mu_B$

- c.  $[\text{FeCl}_4]^-$   $n = 5$   
 $\mu = \sqrt{5(7)} = 5.9 \mu_B$

- d.  $[\text{Fe}(\text{CN})_6]^{2-}$   $n = 1$

e.  $[\text{Ni}(\text{H}_2\text{O})_6]^{2+}$   $n = 2$   
 $\mu = \sqrt{2(4)} = 2.8 \mu_B$

f.  $[\text{Cu}(\text{en})_2(\text{H}_2\text{O})_2]^{2+}$   $n = 1$   
 $\mu = \sqrt{1(3)} = 1.7 \mu_B$

$\mu = \sqrt{1(3)} = 1.7 \mu_B$

# Fifth Edition Miessler Answers

**Matt Walker**



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