

## Unit 8 – Stoichiometry Worksheet 1: Mole relationships

For each of the problems below, use the five steps discussed in class. Answer on a separate piece of paper.

1. Write the balanced chemical equation
2. Identify what is given (with units) and what you want to find (with units) in a BCA table
3. Use the mole ratio from the balanced equation to calculate remaining values in the Table (dimensional analysis style). **Show your work beneath the BCA table.**
4. Complete the BCA Table using calculation results
5. Find the desired quantity in your work. Write and box in that answer with unit, substance.

1. Hydrogen sulfide gas, which smells like rotten eggs, burns in air to produce sulfur dioxide and water. How many moles of oxygen gas would be needed to completely burn 8.0 moles of hydrogen sulfide? (The BCA table is done as an example. Make/copy your own on your own piece of paper).

Equation:  $\text{___ H}_{2}\text{S}_{(g)} + \text{___ O}_{2(g)} \rightarrow \text{___ SO}_{2(g)} + \text{___ H}_2\text{O}_{(g)}$

Before:

**FILL IN THE VALUES ON YOUR OWN PAPER**

Change

(SHOW WORK UNDER THE TABLE)

After

2. Propane,  $\text{C}_3\text{H}_8$ , burns in air to form carbon dioxide and water. If 12 moles of carbon dioxide are formed, how many moles of propane were burned?
3. Ammonia,  $\text{NH}_3$ , for fertilizer is produced by causing hydrogen and nitrogen to react at high temperature and pressure. How many moles of ammonia can be made from 0.15 moles of nitrogen gas?
4. The poison gas phosgene,  $\text{COCl}_2$ , reacts with water in the lungs to form hydrochloric acid and carbon dioxide. How many moles of hydrochloric acid would be formed by 0.835 moles of phosgene?
5. Nitric acid,  $\text{HNO}_3$ , is used in the production of trinitrotoluene, (TNT), an explosive. The equation (not balanced) is



- a) How many moles of nitric acid are required to react with 4.5 moles of toluene? (make your own BCA table)
  - b) How many moles of TNT should be produced?
6. Iron metal and oxygen combine to form magnetite, the magnetic oxide of iron,  $\text{Fe}_3\text{O}_4$ .
    - a) How many moles of iron can be converted to magnetite by 8.80 moles of pure oxygen? (make your own BCA table)
    - b) How many moles of the magnetic iron oxide would be produced?

# Modeling Chemistry Unit 8 Mole Relationships Answers

**Marco Cascella**



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## **Table of Contents Modeling Chemistry Unit 8 Mole Relationships Answers**

1. Understanding the eBook Modeling Chemistry Unit 8 Mole Relationships Answers
  - The Rise of Digital Reading Modeling Chemistry Unit 8 Mole Relationships Answers
  - Advantages of eBooks Over Traditional Books
2. Identifying Modeling Chemistry Unit 8 Mole Relationships Answers
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Modeling Chemistry Unit 8 Mole Relationships Answers
  - User-Friendly Interface
4. Exploring eBook Recommendations from Modeling Chemistry Unit 8 Mole Relationships Answers

- Personalized Recommendations
  - Modeling Chemistry Unit 8 Mole Relationships Answers User Reviews and Ratings
  - Modeling Chemistry Unit 8 Mole Relationships Answers and Bestseller Lists
5. Accessing Modeling Chemistry Unit 8 Mole Relationships Answers Free and Paid eBooks
    - Modeling Chemistry Unit 8 Mole Relationships Answers Public Domain eBooks
    - Modeling Chemistry Unit 8 Mole Relationships Answers eBook Subscription Services
    - Modeling Chemistry Unit 8 Mole Relationships Answers Budget-Friendly Options
  6. Navigating Modeling Chemistry Unit 8 Mole Relationships Answers eBook Formats
    - ePub, PDF, MOBI, and More
    - Modeling Chemistry Unit 8 Mole Relationships Answers Compatibility with Devices
    - Modeling Chemistry Unit 8 Mole Relationships Answers Enhanced eBook Features
  7. Enhancing Your Reading Experience
    - Adjustable Fonts and Text Sizes of Modeling Chemistry Unit 8 Mole Relationships Answers
    - Highlighting and Note-Taking Modeling Chemistry Unit 8 Mole Relationships Answers
    - Interactive Elements Modeling Chemistry Unit 8 Mole Relationships Answers
  8. Staying Engaged with Modeling Chemistry Unit 8 Mole Relationships Answers
    - Joining Online Reading Communities
    - Participating in Virtual Book Clubs
    - Following Authors and Publishers Modeling Chemistry Unit 8 Mole Relationships Answers
  9. Balancing eBooks and Physical Books Modeling Chemistry Unit 8 Mole Relationships Answers
    - Benefits of a Digital Library
    - Creating a Diverse Reading Collection Modeling Chemistry Unit 8 Mole Relationships Answers
  10. Overcoming Reading Challenges
    - Dealing with Digital Eye Strain
    - Minimizing Distractions
    - Managing Screen Time
  11. Cultivating a Reading Routine Modeling Chemistry Unit 8 Mole Relationships Answers
    - Setting Reading Goals Modeling Chemistry Unit 8 Mole Relationships Answers
    - Carving Out Dedicated Reading Time
  12. Sourcing Reliable Information of Modeling Chemistry Unit 8 Mole Relationships Answers

- Fact-Checking eBook Content of Modeling Chemistry Unit 8 Mole Relationships Answers
- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
- 14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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