

# DEVOIR DE CONTROLE N°2

Date: 10 -02-2020

Durée : 2H

- le sujet comporte 2 exercices de chimie et 2 exercices de physique.
- On donnera l'expression littérale avant de passer à l'application numérique.
- L'utilisation de la calculatrice scientifique est autorisée.



## CHIMIE :(9 POINTS)

### Exercice n°1 :

Un acide carboxylique (A) de masse molaire  $M = 74 \text{ g.mol}^{-1}$ .

- Déterminer la formule semi développée et le nom de (A).
- On dissout une masse  $m=0,74\text{g}$  de (A) dans l'eau distillée afin de préparer  $100\text{mL}$  d'une solution (S) d'acide de concentration molaire C. Le  $\text{pH}$  de cette solution est égal à 3.
  - Calculer la concentration molaire C.
  - Comparer C et  $[H_3O^+]$ . Ecrire son équation de dissociation dans l'eau.
  - Déduire la molarité des espèces chimiques présentes dans la solution (sauf l'eau).  
On rappelle qu'à  $25^\circ\text{C}$  :  $[H_3O^+][OH^-]=10^{-14}$
- On fait réagir un volume  $V^*=50\text{mL}$  de la solution (S) avec  $m^*=0,1308\text{g}$  de zinc, il se dégage un gaz qui provoque une légère détonation en présence d'une flamme.  
Quel est le gaz dégagé ? Calculer son volume.
- On fait réagir 5 moles de (A) avec 5 moles d'alcool (B) qui donne à la fin de la réaction 3 moles d'un produit (C) formé par 6 atomes de carbone de plus 3 moles d'eau.
  - Déterminer la formule brute de l'alcool (B) ainsi que la formule semi-développé de (C) et son nom.
  - Dégager de ce qui précède la propriété de cette réaction.

0,75B

0,75B

0,75B

0,75B

1C

1C

0,5B

### Exercice n°2 :

La combustion d'une masse  $m_1=2\text{g}$  d'un alcool (A)  $C_nH_{2n+2}O$  fournit  $m_2=5\text{g}$  de dioxyde de carbone.

- Ecrire l'équation générale de la réaction de combustion.
- Montrer que la masse molaire de l'alcool (A) peut s'écrire  $M=17,6n$ .
- Déduire alors la formule brute de (A)
- L'oxydation ménagée de (A) par le bichromate de potassium ( $2K^+$ ,  $Cr_2O_7^{2-}$ ) en milieu acide donne un composé (B) qui réagit avec le 2,4 DNPH mais ne réagit pas avec le réactif de Schiff.
  - Préciser la fonction chimique de (B) ?
  - Déduire alors tous les isomères possibles de (A).
- Ecrire l'équation de la réaction d'oxydation ménagée en utilisant les formules brutes.
- La déshydratation intramoléculaire de (A) donne un seul composé organique (C). Identifier les composés organiques (A), (B) et (C).

0,5A

0,5 C

0,5B

0,25A

0,75A

1C

1C

On donne  $\text{g. mol}^{-1}$  :  $M(C)=12$  ;  $M(O)=16$  ;  $M(H)=1$  ;  $M(Zn)=65,4$  et  $V_m=24 \text{ L.mol}^{-1}$



# Devoir De Controle N 2 3 M Exercice1

**France. Parlement (1946- ). Assemblée  
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