

LA COMPOSIZIONE DELLA MATERIA

L'oggetto di studio della chimica è la **materia**. La chimica è infatti la disciplina che studia:

- La struttura e la composizione della materia
- Le trasformazioni che la materia subisce
- L'energia coinvolta in queste trasformazioni

MISCELE E SOSTANZE

La materia è composta di **atomi**, particelle piccolissime che possono combinarsi tra loro in modo diverso.

Riguardo alla composizione chimica, la materia può essere sotto forma di:

- **Miscela** (o miscuglio)
- **Sostanza** (o sostanze pure)

MISCELE

Una miscela è un insieme fisico di due o più sostanze e ha una composizione variabile.

In base alla possibilità di distinguerne i componenti, le miscele vengono divise in:

- **Eterogenee**: miscele formate da due o più fasi (una **fase** è una porzione di un sistema in cui le proprietà fisiche intensive sono identiche in ogni punto, separata dal resto da superfici limitanti) che non hanno composizione uniforme in tutti i punti del campione e che i componenti sono facilmente distinguibili.
Esempi di miscele eterogenee sono le rocce, il sangue e il latte.
- **Omogenee**: miscele formate da una sola fase che hanno una composizione uniforme in tutti i punti del campione e che i componenti non sono distinguibili. Le miscele omogenee sono dette anche **soluzioni**.
Esempi di miscele omogenee sono l'acqua di mare, il vino e l'aria.

Separando una miscela nei suoi componenti con metodi fisici (es. filtrazione o distillazione) si ottengono sostanze pure. La scelta del metodo di separazione dipende dallo stato fisico dei componenti e dalle loro proprietà chimiche e fisiche.

I **METODI DI SEPARAZIONE** maggiormente utilizzati sono:

- **Filtrazione**: permette di separare i componenti di una miscela eterogenea costituita da un liquido e un solido sfruttando la diversa dimensione delle particelle. Essa si realizza facendo passare la miscela attraverso un filtro di carta assorbente.
- **Distillazione**: permette di separare i componenti di una miscela eterogenea sfruttandone la diversa temperatura di ebollizione. Essa può essere:

Semplice quando si separa un liquido da una soluzione solido-liquido

Frazionata quando si separa un liquido da una soluzione di due o più liquidi

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D Siedentop



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