

# AP Chemistry: Bonding

## Multiple Choice

41. Which of the following molecules has the shortest bond length?

- (A)  $N_2$  (B)  $O_2$  (C)  $Cl_2$  (D)  $Br_2$  (E)  $I_2$

51.  $\pi$  bonding occurs in each of the following species EXCEPT...

- (A)  $CO_2$  (B)  $C_2H_4$  (C)  $CN^-$  (D)  $C_6H_6$  (E)  $CH_4$

60. Which of the following has a zero dipole moment?

- (A)  $HCN$  (B)  $NH_3$  (C)  $SO_2$  (D)  $NO_2$  (E)  $PF_5$

8. Use the following answers for questions 8 - 9.

- (A) A network solid with covalent bonding  
(B) A molecular solid with zero dipole moment  
(C) A molecular solid with hydrogen bonding  
(D) An ionic solid  
(E) A metallic solid

8. Solid ethyl alcohol,  $C_2H_5OH$

9. Silicon dioxide,  $SiO_2$

80. For which of the following molecules are resonance structures necessary to describe the bonding satisfactorily?

- (A)  $H_2S$  (B)  $SO_2$  (C)  $CO_2$  (D)  $OF_2$  (E)  $PF_3$

Hydrogen Halide	Normal Boiling Point, $^{\circ}C$
HF	19
HCl	-85
HBr	-67
HI	-35

18. The liquefied hydrogen halides have the normal boiling points given above. The relatively high boiling point of HF can be correctly explained by which of the following?

- (A) HF gas is more ideal. (C) HF molecules have a smaller dipole moment.  
(B) HF is the strongest acid. (D) HF is much less soluble in water.  
(E) HF molecules tend to form hydrogen bonds.

42. The  $SbCl_5$  molecule has trigonal bipyramid structure. Therefore, the hybridization of Sb orbitals should be...

- (A)  $sp^2$  (B)  $sp^3$  (C)  $sp^2d$  (D)  $sp^3d$  (E)  $sp^3d^2$

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