

Name _____

SHORT ANSWER. Write the word or phrase that best completes each statement or answers the question.

- 1) What is the solvent in an aqueous solution?

ESSAY. Write your answer in the space provided or on a separate sheet of paper.

- 2) Explain why water is a good solvent for ionic substances.

MULTIPLE CHOICE. Choose the one alternative that best completes the statement or answers the question.

- 3) Of the species below, only _____ is
- not
- an electrolyte.

A) HCl B) Ar C) NaCl D) KOH E) Rb_2SO_4

- 4) What is the molarity of an aqueous solution containing 75.3 g of glucose (
- $\text{C}_6\text{H}_{12}\text{O}_6$
-) in 35.5 mL of solution?

A) 0.197 B) 2.12 C) 3.52 D) 1.85 E) 11.8

- 5) The concentration of species in 500 mL of a 2.104 M solution of sodium sulfate is _____ M sodium ion and _____ M sulfate ion.

A) 4.208, 2.104 B) 1.052, 1.052 C) 2.104, 1.052 D) 2.104, 4.208 E) 2.104, 2.104

- 6) What is the molarity of a solution prepared by diluting 43.72 mL of 1.005 M aqueous
- $\text{K}_2\text{Cr}_2\text{O}_7$
- to 500 mL?

A) 0.879 B) 87.9 C) 0.0879 D) 0.0115 E) 0.0218

- 7) When aqueous solutions of _____ are mixed, a precipitate forms.

A) NaI and KBr
B) Li_2CO_3 and CsI
C) K_2SO_4 and CrCl_3
D) NiBr_2 and AgNO_3
E) KOH and $\text{Ba}(\text{NO}_3)_2$

- 8) What are the spectator ions in the reaction between KOH (aq) and
- HNO_3
- (aq)?

A) H^+ and NO_3^-
B) OH^- only
C) K^+ and NO_3^-
D) K^+ and H^+
E) H^+ and OH^- **SHORT ANSWER.** Write the word or phrase that best completes each statement or answers the question.

- 9) What is the concentration of HCl in a 0.100 M solution of HCl? Why?

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