

Name _____

MULTIPLE CHOICE. Choose the one alternative that best completes the statement or answers the question.

- 1) A chemical reaction that absorbs heat from the surroundings is said to be _____ and has a _____ ΔH at constant pressure.
- endothermic, positive
 - endothermic, negative
 - exothermic, negative
 - exothermic, positive
 - exothermic, neutral

- 2) The reaction



is _____, and therefore heat is _____ by the reaction.

- exothermic, released
 - exothermic, absorbed
 - endothermic, released
 - endothermic, absorbed
 - thermo-neutral, neither released nor absorbed
- 3) The value of ΔH° for the reaction below is -126 kJ . How much heat (in kJ) is released when 2.00 mol of NaOH is formed in the reaction?



- A) 252 B) -126 C) 7.8 D) 63 E) 3.9

- 4) The value of ΔH° for the reaction below is -790 kJ . The enthalpy change accompanying the reaction of 0.95 g of S is _____ kJ.



- A) -23 B) 12 C) -12 D) 23 E) -790

- 5) The value of ΔH° for the reaction below is -6535 kJ . How many kJ of heat are released in the combustion of 16.0 g of $\text{C}_6\text{H}_6\text{(l)}$?



- A) 673 B) 2.68×10^3 C) 5.23×10^4 D) -6535 E) 1.34×10^3

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