

Name _____

MULTIPLE CHOICE. Choose the one alternative that best completes the statement or answers the question.

- 1) A chemical reaction that absorbs heat from the surroundings is said to be _____ and has a _____ ΔH at constant pressure.
- A) endothermic, positive
 - B) endothermic, negative
 - C) exothermic, negative
 - D) exothermic, positive
 - E) exothermic, neutral

- 2) The reaction



is _____, and therefore heat is _____ by the reaction.

- A) exothermic, released
 - B) exothermic, absorbed
 - C) endothermic, released
 - D) endothermic, absorbed
 - E) thermoneutral, neither released nor absorbed
- 3) The value of ΔH° for the reaction below is -126 kJ. How much heat (in kJ) is released when 2.00 mol of NaOH is formed in the reaction?



- A) 252
 - B) -126
 - C) 7.8
 - D) 63
 - E) 3.9
- 4) The value of ΔH° for the reaction below is -790 kJ. The enthalpy change accompanying the reaction of 0.95 g of S is _____ kJ.



- A) -23
 - B) 12
 - C) -12
 - D) 23
 - E) -790
- 5) The value of ΔH° for the reaction below is -6535 kJ. How many kJ of heat are released in the combustion of 16.0 g of C_6H_6 (l)?



- A) 673
- B) 2.68×10^3
- C) 5.23×10^4
- D) -6535
- E) 1.34×10^3

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Victor M. Corman



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