

AP Chemistry: Chapter 5-6 Student Notes

Objectives

Chapter 5: Gases

- ☐ Kinetic Molecular Theory of Gases
- ☐ Understand pressure and its units
- ☐ Boyles-Charles-Gay-Lussac-Avogadro's Laws (Conceptual and Calculations)
- ☐ Ideal Gas Law
- ☐ Gas Stoichiometry
- ☐ Daltons Law of Partial Pressures (and applications to Gas Stoichiometry)
- ☐ Grahams Law
- ☐ Real verses Ideal Gases

Chapter 6: Thermochemistry

- ☐ The First Law of Thermodynamics
- ☐ Energy in a chemical reaction (exothermic and endothermic)
- ☐ Enthalpy ΔH (4 ways to calculate it)
- ☐ Finding ΔH using calorimetry
- ☐ Hess' Law version 1: Adding reactions
- ☐ Hess' Law version 2: $\sum \text{Products} - \sum \text{Reactants}$
- ☐ Different kinds of ΔH

Kinetic Molecular Theory of Gases

Kinetic = _____

Molecular = _____

So the KMT is the _____ theory

All molecules _____

Solids move _____

Liquids _____

Gases move in _____ - _____ motion

How fast do they move?

Use $KE = \frac{1}{2}mv^2$: Solve for velocity

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macam macam pola lantai lengkap dengan contohnya pahami - Oct 05 2022

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pola lantai tersebut memiliki tujuan dan makna yang berbeda tari nusantara yang pola lantainya segi empat diantaranya yaitu pada dasarnya pola lantai jenis ini bertujuan untuk membantu penari bergerak

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