

# STUDY CARD

## For CAMPBELL/RECE BIOLOGY Seventh Edition

### Exploring Life (Ch. 1)

1. Life exists from the microscopic scale up to the global.
2. Living systems are much more than the sum of their parts.
3. Organisms include a great diversity of species.
4. Evolution accounts for life's unity and diversity.
5. Biologists use discovery science and hypothesis-based science to explore life.

### The Chemical Context of Life (Ch. 2)

1. Matter consists of chemical elements in pure form and in combinations called compounds.
2. An element's properties depend on its structure:
  - a. Protons and neutrons are in the nucleus.
  - b. Electrons exist at different energy levels in electron shells; the outermost is the valence shell.
3. Atoms with incomplete valence shells can form chemical bonds with other such atoms by sharing, gaining, or losing electrons.
  - a. In a covalent bond, two atoms share a pair of electrons. (Double or triple bonds can also form.)
  - b. Electrons transfer from one atom to another, based on the attraction between oppositely charged ions.
  - c. Weak bonds include hydrogen bonds and van der Waals interactions.
4. Chemical reactions make and break chemical bonds as reactants are converted to products.

### Water (Ch. 3)

1. Water molecules are polar. Hydrogen bonding between water molecules gives water unique properties that allow life to exist on Earth.
  - a. Cohesion and adhesion allow capillary action, which functions during water transport in plants.
  - b. Water's high specific heat and high heat of vaporization moderate temperatures on Earth.
  - c. Ice floats because hydrogen bonds lock water into a crystalline lattice that reduces its density.



Figure 3.23 (a) and (b)

- a. Water is a versatile solvent.
  - b. Hydrophobic substances have an affinity for one another, while hydrophilic substances repel water.
2. Dissociation of water molecules affects the pH.
    - a. Acids, which have a pH  $< 7$ .
    - b. Bases, which have a pH  $> 7$ .
    - c. Buffers minimize changes in pH.

### Carbon and Molecular Diversity (Ch. 4)

1. Organic chemistry is the study of carbon compounds.
2. The four atoms that are often found in molecules forming bonds with up to four other atoms.
  - a. Carbon atoms form the skeletons of most organic molecules. The carbon atom (C), if only, can form four bonds.
  - b. Boron has the same number of atoms and elements but different structures and properties.
  - c. These four groups are the parts of molecules involved in chemical reactions.

3. ATP is an important energy-transforming molecule:  $ADP + \text{inorganic } P \rightarrow \text{ATP} + \text{energy}$ .

### Macromolecules (Ch. 5)

1. Most macromolecules are polymers synthesized from monomers by the dehydration of water.
  - a. Polymers are characterized by monomers (or addition of a water molecule) (hydrolysis).
  - b. Variation in the linear sequence of monomers results in the diversity of polymers.
2. Four major classes of biological macromolecules exist and occur in cells.
  - a. Carbohydrates include sugars (monosaccharides) and sugar polymers (polysaccharides).
    - i. Most monosaccharides are multiples of  $\text{CH}_2\text{O}$ .
    - ii. Starch (in plants) and glycogen (in animals) are glucose polymers that store energy.
    - iii. Cellulose (in plants) and chitin (in animals) are polysaccharides that provide structural support.
  - b. Lipids are hydrophobic.
    - i. Fats (glycerol linked to three fatty acid chains) can be saturated or unsaturated; they store energy.
    - ii. Phospholipids (glycerol, two fatty acids, and phosphate) have hydrophobic and hydrophilic regions; they form a bilayer in water.
    - iii. Steroids (a carbon skeleton with four fused rings) include cholesterol and sex hormones.
  - c. Proteins have diverse structures and functions.
    - i. A protein consists of one or more polypeptides, each a polymer of amino acids (each a set of  $\text{C}_\alpha$ ).
    - ii. A peptide bond joins the carbonyl end of one amino acid with the amino group of the next.



Figure 5.1

3. Four levels of protein structure lead to the proper shape and function of each protein.
  - a. Nucleic acids (DNA and RNA) store and pass on hereditary information.
    - i. DNA and RNA are polymers of nucleotides.
    - ii. A nucleotide has three parts: a 5-carbon sugar, a nitrogenous base, and a phosphate group.
    - iii. DNA has deoxyribose as the sugar; RNA has ribose.
    - iv. DNA has A, C, G, and T as bases; RNA has A, C, G, and U.
    - v. RNA is a double-stranded helix; RNA is single-stranded.
    - vi. A gene (hereditary unit) is a stretch of DNA that is transcribed into RNA, which is then translated into the amino acid sequence of a polypeptide.

### A Tour of the Cell (Ch. 6)

1. All organisms are composed of one or more cells.
2. Cells are bounded by a plasma membrane and contain cytosol, organelles, and one or more chromosomes, and many ribosomes.

3. Prokaryotic cells usually have a cell wall but lack a nucleus and other membrane-bounded organelles.



Figure 6.1

4. Eukaryotic cells are generally larger than prokaryotic cells and have a nucleus and membrane-bounded organelles.
  - a. An animal cell.

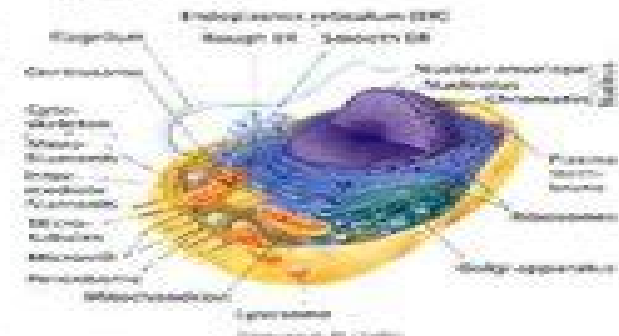


Figure 6.2 (left)

- a. A plant cell.

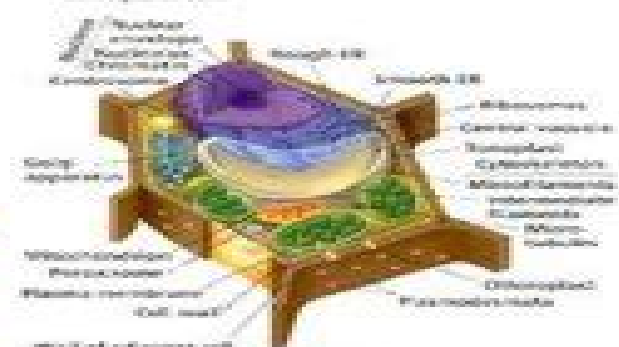


Figure 6.2 (right)

5. Within the nucleus, chromosomes made up of chromatin (DNA and proteins) carry the genetic instructions for making polypeptides.
6. Mitochondria are the organelles in which polypeptides are synthesized.
7. The endoplasmic reticulum (ER) is a network of membranes.
  - a. Smooth ER functions in lipid synthesis, carbohydrate metabolism, and detoxification.
  - b. Rough ER has ribosomes that synthesize secretory proteins, which enter the ER lumen. Rough ER also synthesizes membrane proteins and phospholipids.

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**Ali Murat SEDEF**



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