



Periodic Table of the Elements

Legend:

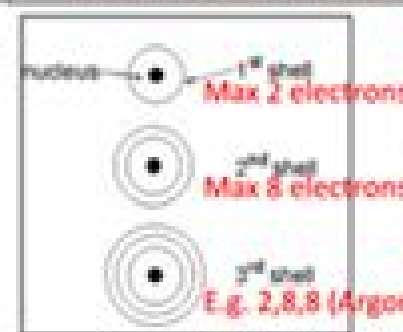
- metals
- non-metals
- metalloids
- transition metals
- alkali metals
- alkali earth metals
- transition metals
- lanthanoids
- actinoids

An atom:
Number of Protons = Number of Electrons
Mass number = Protons + Neutrons

C2 Chemistry

Periodic table arranges elements by
ATOMIC NUMBER (proton number/the small one)

The number of outer shell electrons match the group the element is found in.
E.g. Lithium 2,1 is a group 1 element.

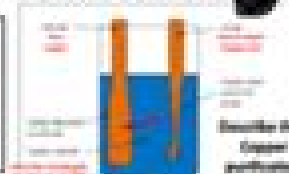


Protons and neutrons both have a mass of 1

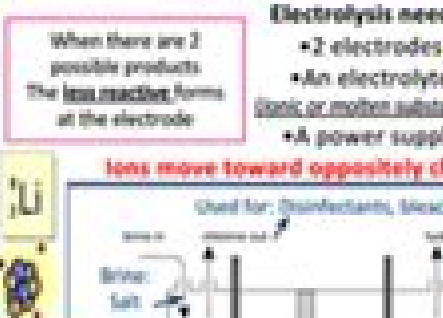
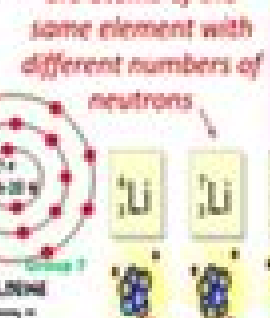
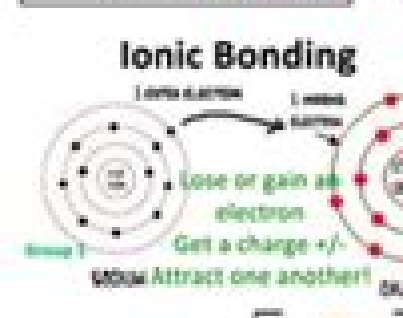
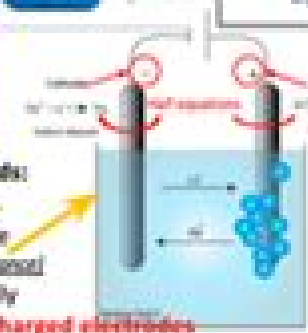
ISOTOPES
are atoms of the same element with different numbers of neutrons

OIL RIG

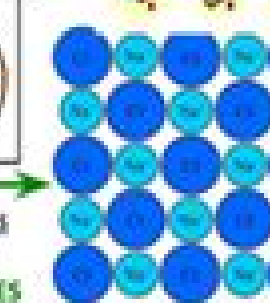
Oxidation is loss (of electrons)
Reduction is gain (of electrons)



Electrolysis:
Splitting up a substance using electricity



Giant Lattice Structure
Always metals with non-metals
+ ION and - ION attracted by **STRONG ELECTROSTATIC FORCES**



Electrolysis of Brine
Used for: Bleach, paper and soap
As a solid: High melting and Boiling Point
Conducts electricity when melted or dissolved: ions move freely and carry a charge

Electrolysis of Brine
Used for: Bleach, paper and soap
As a solid: High melting and Boiling Point
Conducts electricity when melted or dissolved: ions move freely and carry a charge

NANO-TECHNOLOGY
Means REALLY REALLY REALLY small!!!!

Aqa Gcse Chemistry Revision Notes

Jan Hayes

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author's name.

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