

AP Chemistry Chapter 8 Answers – Zumdahl

8.87

The first step always is to draw a valid Lewis structure when predicting molecular structure. When resonance is possible, only one of the possible resonance structures is necessary to predict the correct structure because all resonance structures give the same structure. The Lewis structures are in Exercises 8.67 and 8.73. The structures and bond angles for each follow.

- 8.67 a. HCN: linear, 180°
 b. PH_3 : trigonal pyramid, $< 109.5^\circ$
 c. CHCl_3 : tetrahedral, 109.5°
 d. NH_4^+ : tetrahedral, 109.5°
 e. H_2CO : trigonal planar, 120°
 f. SeF_2 : V-shaped or bent, $< 109.5^\circ$
 g. CO_2 : linear, 180°

h and i. O_2 and HBr are both linear, but there is no bond angle in either.

Note: PH_3 and SeF_2 both have lone pairs of electrons on the central atom which result in bond angles that are something less than predicted from a tetrahedral arrangement (109.5°). However, we cannot predict the exact number. For the solutions manual, we will insert a less than sign to indicate this phenomenon. For bond angles equal to 120° , the lone pair phenomenon isn't as significant as compared to smaller bond angles. For these molecules, e.g., NO_2^- , we will insert an approximate sign in front of the 120° to note that there may be a slight distortion from the VSEPR bond angle.

- 8.73 a. NO_2^- : V-shaped, approximately 120° ; NO_3^- : trigonal planar, 120°
 N_2O_4 : trigonal planar, 120° about both N atoms
 b. OCN^- , SCN^- , and N_3^- are all linear with 180° bond angles.

8.89

From the Lewis structures (see Exercise 8.71), Br_3^- would have a linear molecular structure, ClF_3 would have a T-shaped molecular structure and SF_4 would have a see-saw molecular structure.

For example, consider ClF_3 (28 valence electrons): The central Cl atom is surrounded by 5 electron pairs, which requires a trigonal bipyramidal geometry. Since there are 3 bonded atoms and 2 lone pairs of electrons about Cl, we describe the molecular structure of ClF_3 as T-shaped with predicted bond angles of about 90° . The actual bond angles would be slightly less than 90° due to the strong repulsive effect of the lone pair electrons as compared to the bonding electrons.

8.91

- a. SeO_3 has a trigonal planar molecular structure with all bond angles equal to 120° . Note that any one of the resonance structures could be used to predict molecular structure and bond angles.

Ap Chemistry Chapter 8 Answers Zumdahl 8

Steven S. Zumdahl, Kenneth C. Brooks



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Chemistry Principles Student Solutions Guide Steven S. Zumdahl, 2001-08 End of chapter problems with blue numbers appear in this guide These answers also appear at the end of the text These solutions include a mix of odd and even numbered problems comprising about half of the total problem sets found in the text **Science Books & Films**, 1987

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1. Understanding the eBook Ap Chemistry Chapter 8 Answers Zumdahl 8
 - The Rise of Digital Reading Ap Chemistry Chapter 8 Answers Zumdahl 8
 - Advantages of eBooks Over Traditional Books
2. Identifying Ap Chemistry Chapter 8 Answers Zumdahl 8
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Ap Chemistry Chapter 8 Answers Zumdahl 8
 - User-Friendly Interface
4. Exploring eBook Recommendations from Ap Chemistry Chapter 8 Answers Zumdahl 8
 - Personalized Recommendations
 - Ap Chemistry Chapter 8 Answers Zumdahl 8 User Reviews and Ratings
 - Ap Chemistry Chapter 8 Answers Zumdahl 8 and Bestseller Lists
5. Accessing Ap Chemistry Chapter 8 Answers Zumdahl 8 Free and Paid eBooks

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 - Ap Chemistry Chapter 8 Answers Zumdahl 8 Budget-Friendly Options
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 - ePub, PDF, MOBI, and More
 - Ap Chemistry Chapter 8 Answers Zumdahl 8 Compatibility with Devices
 - Ap Chemistry Chapter 8 Answers Zumdahl 8 Enhanced eBook Features
 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Ap Chemistry Chapter 8 Answers Zumdahl 8
 - Highlighting and Note-Taking Ap Chemistry Chapter 8 Answers Zumdahl 8
 - Interactive Elements Ap Chemistry Chapter 8 Answers Zumdahl 8
 8. Staying Engaged with Ap Chemistry Chapter 8 Answers Zumdahl 8
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Ap Chemistry Chapter 8 Answers Zumdahl 8
 9. Balancing eBooks and Physical Books Ap Chemistry Chapter 8 Answers Zumdahl 8
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Ap Chemistry Chapter 8 Answers Zumdahl 8
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine Ap Chemistry Chapter 8 Answers Zumdahl 8
 - Setting Reading Goals Ap Chemistry Chapter 8 Answers Zumdahl 8
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of Ap Chemistry Chapter 8 Answers Zumdahl 8
 - Fact-Checking eBook Content of Ap Chemistry Chapter 8 Answers Zumdahl 8
 - Distinguishing Credible Sources
 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development

- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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