

Chemical Reactions: Oxidation-Reduction Reactions

1. All oxidation-reduction reactions involve a transfer of electrons between atoms or groups of atoms. Which reaction or set of steps below does not show either a transfer of electrons or a transfer of atoms?
- (A) $\text{Al} + \text{Cl}_2 \rightarrow \text{AlCl}_3$
 - (B) $\text{H}_2 + \text{O}_2 \rightarrow \text{H}_2\text{O}$
 - (C) $\text{Na} + \text{H}_2\text{O} \rightarrow \text{NaOH} + \text{H}_2$
 - (D) $\text{Ca} + \text{O}_2 \rightarrow \text{CaO}$
- Answer: (D)
2. For which of the following substances would the higher oxidation state of phosphorus indicate that it has been oxidized?
- (A) PO_4^{3-} and PO_4^2-
 - (B) H_3PO_4 and H_2PO_4^-
 - (C) H_3PO_3 and H_2PO_3^-
 - (D) H_3PO_2 and H_2PO_2^-
- Answer: (D)
3. Phosphorus undergoes its most common oxidation state changes in the following sequence.
- (A) P_4 to P_2O_5 to P_2O_{10}
 - (B) P_2O_5 to P_2O_{10} to P_4
 - (C) P_4 to P_2O_{10} to P_2O_5
 - (D) P_2O_{10} to P_2O_5 to P_4
- Answer: (A)
4. For some acids, the chlorine atom takes part in
- (A) conversion of water to hydrochloric acid
 - (B) oxygenated oxychlorides
 - (C) conversion of chlorides to chlorinated oxyacids
 - (D) reduction of other chlorinated oxyacids to chlorinated oxyacids
- Answer: (A)
5. In which of the following reactions does iodine act as both an oxidizing agent and a reducing agent?
- (A) $\text{I}_2 + \text{H}_2\text{O} \rightarrow \text{HIO}_3 + \text{HI}$
 - (B) $\text{I}_2 + \text{H}_2\text{S} \rightarrow \text{H}_2\text{S}_2\text{O}_3 + \text{HI}$
 - (C) $\text{I}_2 + \text{H}_2\text{O} \rightarrow \text{HIO}_3 + \text{H}_2$
 - (D) $\text{I}_2 + \text{H}_2\text{O} \rightarrow \text{HIO}_3 + \text{H}_2\text{I}$
- Answer: (C)
6. In an exothermic reaction, which substance is oxidized? (An exothermic reaction releases heat energy.)
- (A) $\text{Zn} + \text{H}_2\text{O} \rightarrow \text{Zn(OH)}_2 + \text{H}_2$
 - (B) $\text{Zn} + \text{H}_2\text{S} \rightarrow \text{ZnS} + \text{H}_2$
 - (C) $\text{Zn} + \text{H}_2\text{O} \rightarrow \text{Zn(OH)}_3 + \text{H}_2$
 - (D) $\text{Zn} + \text{H}_2\text{O} \rightarrow \text{Zn(OH)}_4 + \text{H}_2$
- Answer: (D)
7. In an exothermic reaction, which one of the following elements that form binary molecules has the greatest tendency to gain electrons?

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**Matthew T. Provencher,Brian J.
Cole,Anthony A. Romeo,Pascal
Boileau,Nikhil Verma**

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