

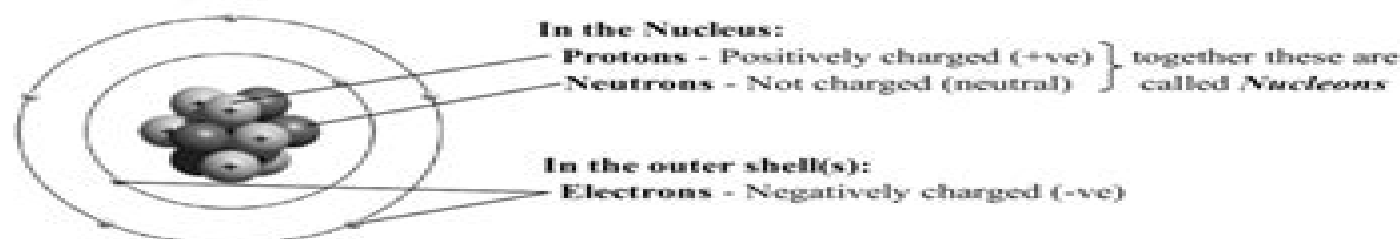
**(1.9) Atomic Structure**

All substances are made up of *atoms* - atoms are very small -  $1/1,000,000,000$  cm (one billionth)

Substances with only *one* type of atom are *elements*

**However:** Atoms are *not* the smallest particles

Atoms are actually made up of *sub-atomic particles* arranged in a *structure* with a *nucleus* and an *outer shell*

**LEARN THIS:****(1.10 / 1.11 / 2.4) Atomic Number, Mass Number, Relative Atomic Mass & Isotopes:**

- Atoms are usually neutral - the number of *protons* (+ve) is equal to the number of *electrons* (-ve)
- The atomic number (or *proton number*) of any element is the number of *protons* in the nucleus of the element - it is *fixed* and *unique* - only *one* element has *that* atomic number
- The mass number (or *nucleon number*) of a particular atom of an element is equal to the number of *protons* plus the number of *neutrons* in the nucleus (i.e. the total number of *nucleons*) - this may vary in different isotopes of the element (*see below*)
- The number of *neutrons* in the nucleus is *roughly* equal to the number of protons but *this may vary* to produce different isotopes of the same element
- The relative atomic mass (RAM or Ar) of any element is the *weighted average mass of the isotopes of the element* - based on the *carbon-12 scale*, where the *isotope carbon-12* has a mass of exactly 12 units (calculations of Ar are based on the *percentages* of different isotopes in any given sample of the element - *see below* (*see also: Edexcel Chemistry p177*))
- Electrons orbit at a relatively *large distance* from the nucleus (atoms are mainly empty space) - their mass is roughly  $1/2000$  ( $1/1836$ ) that of a proton (*2,000 electrons have the same mass as 1 proton*) the mass of the electrons is therefore negligible and can be ignored
- The *electrons* are arranged in concentric *shells* orbiting the nucleus:-
  - The *inner shell* can contain a *maximum* of *two* electrons
  - The *outer shells* can contain a *maximum* of *eight* electrons
 } *this 2, 8, 8 rule only applies to the first 20 elements*
- (2.4) The number of *electrons in the outer shell (Valence Electrons)* gives the element's *group* number (vertical column) and *explains why elements in the same group have similar properties*
- The number of *shells* orbiting the nucleus gives the element's *period* number (horizontal row)
- (1.13 / 2.1) The discovery of this atomic structure of the elements allows them to be arranged in the **Periodic Table**, where the vertical columns are **Groups** and the horizontal rows are **Periods**

# As Chemistry Revision Notes Unit 1 Atomic Structure

**Lei Shi**



## **As Chemistry Revision Notes Unit 1 Atomic Structure:**

## Unveiling the Magic of Words: A Report on "**As Chemistry Revision Notes Unit 1 Atomic Structure**"

In a world defined by information and interconnectivity, the enchanting power of words has acquired unparalleled significance. Their ability to kindle emotions, provoke contemplation, and ignite transformative change is truly awe-inspiring. Enter the realm of "**As Chemistry Revision Notes Unit 1 Atomic Structure**," a mesmerizing literary masterpiece penned by way of a distinguished author, guiding readers on a profound journey to unravel the secrets and potential hidden within every word. In this critique, we shall delve in to the book is central themes, examine its distinctive writing style, and assess its profound affect the souls of its readers.

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