

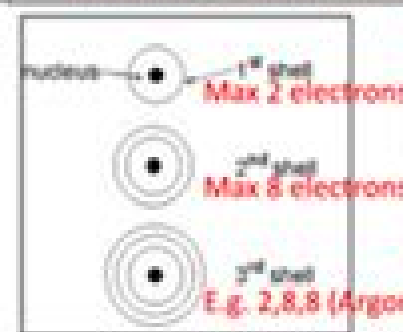
Periodic Table of the Elements

An atom:  
Number of Protons = Number of Electrons  
Mass number = Protons + Neutrons

# C2 Chemistry

Periodic table arranges elements by  
**ATOMIC NUMBER** (proton number/the small one)

The number of outer shell electrons match the group the element is found in.  
E.g. Lithium 2,1 is a group 1 element.

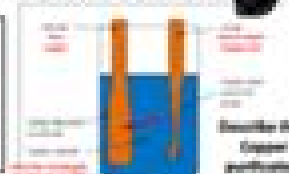


Protons and neutrons both have a mass of 1

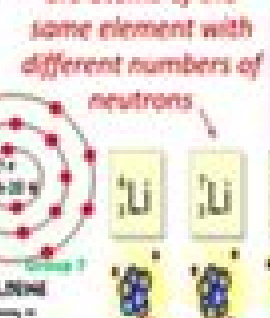
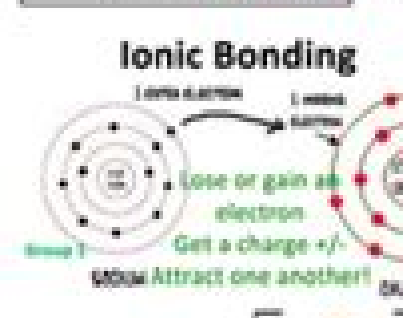
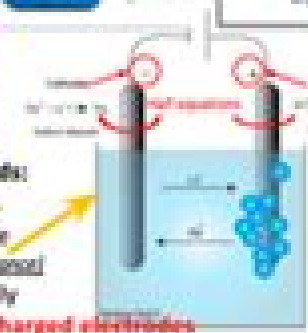
**ISOTOPES**  
are atoms of the same element with different numbers of neutrons

**OIL RIG**

Oxidation is loss (of electrons)  
Reduction is gain (of electrons)



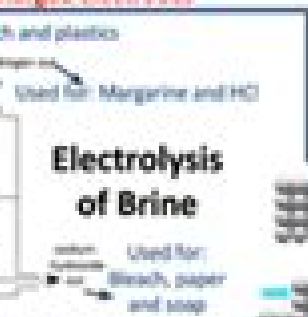
**Electrolysis:**  
Splitting up a substance using electricity



When there are 2 possible products  
The less reactive forms at the electrode

Electrolysis needs:

- 2 electrodes
- An electrolyte (ionic or molten substance)
- A power supply



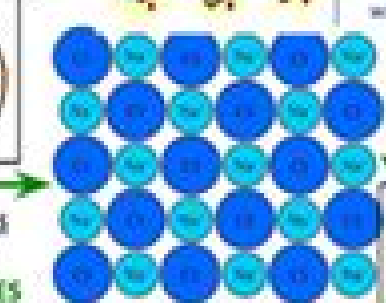
Metal ions only arranged in layers  
**ELECTROSTATIC ATTRACTION**  
between  
**IONS and ELECTRONS**

Metal ions slide over one another making them **MALLEABLE**

Metals conduct **HEAT** well

**Giant Lattice Structure**

Always metals with non-metals  
+ ION and - ION attracted by **STRONG ELECTROSTATIC FORCES**



As a solid: High melting and Boiling Point  
Conducts electricity when melted or dissolved:  
ions move freely and carry a charge

**NANO-TECHNOLOGY**  
Means **REALLY REALLY REALLY** small!!!!

# Aqa Gcse Chemistry Revision

**Christine Horbury**

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author's name.

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