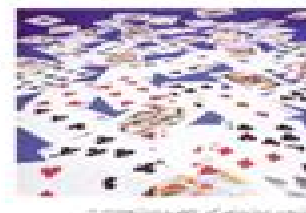




WHAT DRIVES A REACTION TO BE THERMODYNAMICALLY FAVORABLE?

ENTHALPY (ΔH) – heat exchange (exothermic reactions are generally favored)

ENTROPY (ΔS) – dispersal (disorder) of the *matter and energy* of a system (more dispersal/disorder is favored) Nature tends toward chaos! Think about your room at the end of the week. Your mom knows this concept all too well.



Thermodynamically favored processes or reactions are those that involve both a decrease in the internal energy of the components ($\Delta H^\circ < 0$) and an increase in entropy of the components ($\Delta S^\circ > 0$). These processes are necessarily “*thermodynamically favored*” ($\Delta G^\circ < 0$) or negative. If you are using an older textbook, you may see these reactions referred to as “spontaneous”. Avoid that language as you answer free-response questions!

Thermodynamically favored processes occur without outside intervention once the energy of activation has been reached. Thermodynamics does not predict nor take into account the rate of the reaction. That is the realm of kinetics. Some reactions are very fast (like combustion of hydrogen) other reactions are very slow (like graphite turning to diamond) but both are thermodynamically favored.

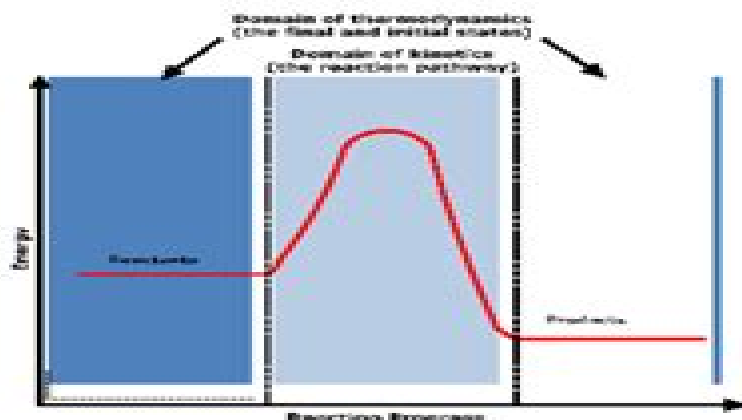
The first law of thermodynamics: Energy can never be created nor destroyed. Therefore, the energy of the universe is constant. This is simply a statement of the law of conservation of energy you’ve know about for quite some time.

The second law of thermodynamics: the universe is constantly increasing the dispersal of matter and energy. Rudolph Clausius “discovered” Δ and gave it its symbol.

The third law of thermodynamics: the entropy of a perfect crystal at 0 K is zero.

[Not a lot of perfect crystals out there so, entropy values are RARELY ever zero—even elements]

So what? This means the absolute entropy of a substance can then be determined at any temp. higher than zero K. (Handy to know if you *ever* need to defend why G & H for elements = 0. . . . BUT S does not!)



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Kingdom and International Baccalaureate in both countries An Englishman he lives in Atlanta Georgia where he teaches at The Westminster Schools He holds a B Sc Hons Chemistry and a Postgraduate Certificate in Education both from the University of Exeter in England In addition to writing this Crash Course Mr Dingle has written The Periodic Table Elements With Style How To Make A Universe With 92 Ingredients and SAT Chemistry Crash Course He is the 2011 winner of the School Library Association of the UK's Information Book Award and in 2012 was honored with the prestigious literary prize Wissenschaftsbuch des Jahre sponsored by the Austrian Ministry of Science and Research **Cholesterol** Anna N.

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