

AP Chemistry Practice Test #1 - Key
Chapters 1, 2, and 3

1. In 1928, 3.3 g of a new element was isolated from 660 kg of the ore molybdenite. The percent by mass of this element in the ore was:
 - a. 5 %
 - b. 6.6 %
 - c. 3.3 %
 - d. 5.0×10^{-4} %
 - e. 2.2 %
2. A quantitative observation
 - a. contains a number and a unit.
 - b. does not contain a number.
 - c. always makes a comparison.
 - d. must be obtained through experimentation.
 - e. none of these
3. Which of the following is the least probable concerning five measurements taken in the lab?
 - a. The measurements are accurate and precise.
 - b. The measurements are accurate but not precise.
 - c. The measurements are precise but not accurate.
 - d. The measurements are neither accurate nor precise.
 - e. All of these are equally probable.
4. A scientist obtains the number 0.045006700 on a calculator. If this number actually has four (4) significant figures, how should it be written?
 - a. 0.4567
 - b. 0.4501
 - c. 0.045
 - d. 0.04500
 - e. 0.04501
5. A piece of indium with a mass of 16.61 g is submerged in 46.3 cm³ of water in a graduated cylinder. The water level increases to 48.6 cm³. The correct value for the density of indium from these data is:
 - a. 7.222 g/cm³
 - b. 7.2 g/cm³
 - c. 0.14 g/cm³
 - d. 0.138 g/cm³
 - e. more than 0.1 g/cm³ away from any of these values.
6. Using the rules of significant figures, calculate the following: 4.0021 – 1.013
 - a. 2.989
 - b. 3
 - c. 2.9891
 - d. 2.99
 - e. 3.0

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