

## Topic: COLLIGATIVE PROPERTIES OF NONELECTROLYTES AND ELECTROLYTE SOLUTIONS

**Colligative properties** – are properties of a solution that depend only on the number and not on the identity of the solute particles.

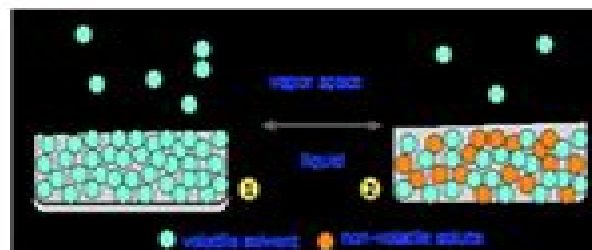
- vapor pressure lowering
- boiling point elevation and
- freezing point depression

### Effect of solute concentration on the colligative properties of solutions

The concentration or amount of nonvolatile solute (i.e., a solute that does not have a vapor pressure of its own) in the solution has an effect on the colligative properties of solutions. The effect would depend on the ratio of the number of particles of solute and solvent in the solution and not on the identity of the solute. However, it is necessary to take into account whether the solute is an electrolyte or a nonelectrolyte.

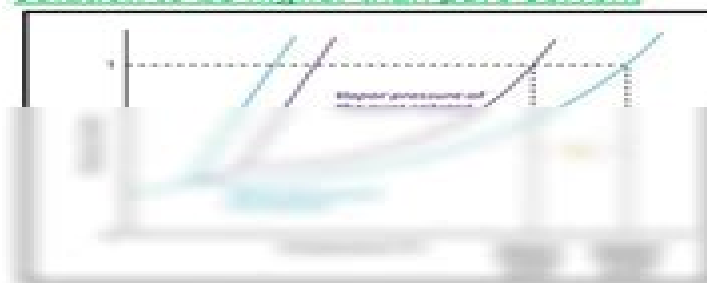
### Effects of electrolyte and nonelectrolyte on colligative properties solutions.

1. **Vapor Pressure Lowering** – The tendency of molecules to leave their liquid state is a dynamic process and equilibrium with the vapor that exists above the liquid. In solution, the presence of solute in the liquid causes the vapor pressure to decrease. The rate at which molecules leave the liquid is reduced, and the rate at which they return to the liquid is increased. This is because the solute particles are in the liquid and are not able to escape as easily as the solvent particles.



2. **Boiling point Elevation** – The addition of a nonvolatile solute lowers the vapor pressure of the solution; consequently the temperature must be raised to restore the vapor pressure of the solution to the value conforming to the pure solvent. Specifically, the temperature at which the vapor pressure is 1 atm will be higher than the normal boiling point by an amount known as the boiling point elevation.

The lowering of the vapor pressure in a solution causes the boiling point of the solution to be higher than pure solvent



Boiling occurs when the vapor pressure of a liquid equals the external pressure. In a solution, the vapor pressure is lower than in the pure solvent, so the boiling point is higher.

# Properties Of Solutions Electrolytes And Nonelectrolytes Lab Report

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