

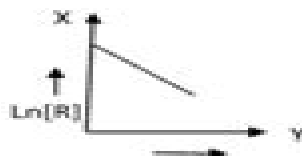
CHEMICAL KINETICS

1 MARK QUESTIONS

1. The decomposition reaction of ammonia gas on platinum surface has a rate constant $K = 2.5 \times 10^{-4} \text{ mol l}^{-1} \text{sec}^{-1}$. What is the order of reaction?
2. For a reaction, $A + B \rightarrow \text{Product}$, the rate law is given by: $r = [A]^2[B]^2$. What is the order of the reaction?
3. What are the units of rate constant for a 1st order reaction?
4. A substance with initial concentration 'a' follows zero order kinetics. In how much time will reaction goes to completion?
5. The reaction $2A + B \rightarrow C + D$ obeys the rate equation, $\text{rate} = K[A]^2[B]$. Find out the order when A is present in large excess.
6. Express the rate of the following reaction in terms of disappearance of hydrogen in the reaction $3H_2 + N_2 \rightarrow 2NH_3$
7. The decomposition reaction of ammonia gas on platinum surface has a rate constant $K = 2.5 \times 10^{-4} \text{ mol l}^{-1} \text{min}^{-1}$. What is the order of reaction?
8. What are the units of rate constant for a 2nd order reaction.
9. In a reaction if the concentration of a reactant A is tripled the rate of reaction becomes twenty-seven times. What is the order of reaction?
10. Identify the reaction order from the following rate constant, $k = 2.3 \times 10^{-5} \text{ L mol}^{-1} \text{ s}^{-1}$.

2 MARKS QUESTIONS

1. Derive the expression for rate constant for zero order reaction.
2. Define the following terms.
 - (i) Molecularity
 - (ii) Rate constant of a reaction.
3. The decomposition of hydrocarbon follows the equation $K = (4.5 \times 10^{11} \text{ sec}^{-1}) e^{-28000K/J}$. Calculate E_a ?
4. For a chemical reaction, $\ln[R]$ vs. time (S) plot is given below. For the reaction
 - (i) What is the order of the reaction?
 - (ii) What is the unit of rate constant K?
 - (iii) Given the relationship between K & t %?
 - (iv) What does the slope of above line indicate?



5. What is meant by pseudo first order reaction? Give an example of pseudo first order reaction and write the rate equation for the same.
6. A reaction is of second order with respect to a reactant. Now how rate is affected if the concentration of the reactant is (i) doubled (ii) reduced to half?
7. (i) For a chemical reaction what is the effect of catalyst on the activation energy?

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1. Understanding the eBook Practice Chemical Kinetics Questions Answer
 - The Rise of Digital Reading Practice Chemical Kinetics Questions Answer
 - Advantages of eBooks Over Traditional Books
2. Identifying Practice Chemical Kinetics Questions Answer
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Practice Chemical Kinetics Questions Answer
 - User-Friendly Interface
4. Exploring eBook Recommendations from Practice Chemical Kinetics Questions Answer
 - Personalized Recommendations
 - Practice Chemical Kinetics Questions Answer User Reviews and Ratings
 - Practice Chemical Kinetics Questions Answer and Bestseller Lists
5. Accessing Practice Chemical Kinetics Questions Answer Free and Paid eBooks

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 - Practice Chemical Kinetics Questions Answer eBook Subscription Services
 - Practice Chemical Kinetics Questions Answer Budget-Friendly Options
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 - Practice Chemical Kinetics Questions Answer Compatibility with Devices
 - Practice Chemical Kinetics Questions Answer Enhanced eBook Features
7. Enhancing Your Reading Experience
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 - Highlighting and Note-Taking Practice Chemical Kinetics Questions Answer
 - Interactive Elements Practice Chemical Kinetics Questions Answer
8. Staying Engaged with Practice Chemical Kinetics Questions Answer
- Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Practice Chemical Kinetics Questions Answer
9. Balancing eBooks and Physical Books Practice Chemical Kinetics Questions Answer
- Benefits of a Digital Library
 - Creating a Diverse Reading Collection Practice Chemical Kinetics Questions Answer
10. Overcoming Reading Challenges
- Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Practice Chemical Kinetics Questions Answer
- Setting Reading Goals Practice Chemical Kinetics Questions Answer
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Practice Chemical Kinetics Questions Answer
- Fact-Checking eBook Content of Practice Chemical Kinetics Questions Answer
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development

- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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