

KINETICS Practice Problems and Solutions

Name: _____
 Period: _____ Date: _____

AP Chemistry
 Dr. Mandes

The following questions represent potential types of quiz questions. Please answer each question completely and thoroughly. The solutions will be posted on-line on Monday.

5. Please do #18 in chapter 12 of your text.

- a. This increases the $[H_2]$, which will increase the rate, but has no effect on k .
- b. Due to Arrhenius equation, changing temperature changes the value of k .
- c. Catalyst function by lowering the activation energy, so due to Arrhenius equation, changing the activation energy changes the value of k .

6. Consider the reaction: $P_4 + 6 H_2 \rightarrow 4 PH_3$. A rate study of this reaction was conducted at 298 K. The data that were obtained are shown in the table.

$[P_4]$, mol/L	$[H_2]$, mol/L	Initial Rate, mol/(L · s)
0.0110	0.0075	3.20×10^{-4}
0.0110	0.0150	6.40×10^{-4}
0.0220	0.0150	6.39×10^{-4}

- a. What is the order with respect to: P_4 0
 H_2 1
- b. Write the rate law for this reaction. $rate = k[H_2]$
- c. Determine the value and units of the rate constant, k . plug and chug using the rate law & data from exp't 1 and solving for k , we get $k = 0.0427 \text{ s}^{-1}$

7. Consider the reaction: $SO_2 + O_3 \rightarrow SO_3 + O_2$. A rate study of this reaction was conducted at 298 K. The data that were obtained are shown in the table.

$[SO_2]$, mol/L	$[O_3]$, mol/L	Initial Rate, mol/(L · s)
0.25	0.40	0.118
0.25	0.20	0.118
0.75	0.20	1.062

- a. What is the order with respect to: SO_2 2
 O_3 0
- b. Write the rate law for this reaction. $rate = k[SO_2]^2[O_3]^0$
- c. Determine the value and units of the rate constant, k . plug and chug using the rate law & data from exp't 1 and solving for k , we get $k = 2.36 \text{ mol/L}^{-2} \text{ s}^{-1}$

8. Consider the following mechanism.

$$A_2 + B_2 \rightarrow R + C \quad (\text{slow})$$

$$A_2 + R \rightarrow C \quad (\text{fast})$$

- a. Write the overall balanced chemical equation. $2 A_2 + B_2 \rightarrow 2 C$
- b. Identify any intermediates within the mechanism. R
- c. What is the order with respect to each reactant? A_2 1st; B_2 1st

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Survival Guide to General Chemistry Patrick E. McMahon, Rosemary McMahon, Bohdan Khomtchouk, 2019-02-13 This work evolved over thirty combined years of teaching general chemistry to a variety of student demographics. The focus is not to recap or review the theoretical concepts well described in the available texts. Instead, the topics and descriptions in this book make available specific detailed step by step methods and procedures for solving the major types of problems in general chemistry. Explanations, instructional process sequences, solved examples, and completely solved practice problems are greatly expanded, containing significantly more detail than can usually be devoted to in a comprehensive text. Many chapters also provide alternative viewpoints as an aid to understanding. Key Features: The authors have included every major topic in the first semester of general chemistry and most major topics from the second semester. Each is written in a specific and detailed step by step process for problem solving, whether mathematical or conceptual. Each topic has greatly expanded examples and solved practice problems containing significantly more detail than found in comprehensive texts. Includes a chapter designed to eliminate confusion concerning acid base reactions, which often persists through working with acid base equilibrium. Many chapters provide alternative viewpoints as an aid to understanding. This book addresses a very real need for a large number of incoming freshmen in STEM fields. **(Free Sample) GO TO Objective NEET Physics Guide with DPP & CPP Sheets**

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87 249 308 and 354 of *Methods in Enzymology* The chapters describe the most essential and widely applied strategies A set of exercises and problems is included to facilitate mastery of these topics The book will aid the reader to design execute and analyze kinetic experiments on enzymes Its emphasis on enzyme inhibition will also make it attractive to pharmacologists and pharmaceutical chemists interested in rational drug design Of the seventeen chapters presented in this new edition ten did not previously appear in the first edition Transient kinetic approaches to enzyme mechanisms Designing initial rate enzyme assay Deriving initial velocity and isotope exchange rate equations Plotting and statistical methods for analyzing rate data Cooperativity in enzyme function Reversible enzyme inhibitors as mechanistic probes Transition state and multisubstrate inhibitors Affinity labeling to probe enzyme structure and function Mechanism based enzyme inactivators Isotope exchange methods for elucidating enzymatic catalysis Kinetic isotope effects in enzyme catalysis Site directed mutagenesis in studies of enzyme catalysis

Reaction Kinetics: Exercises, Programs and Theorems János Tóth, Attila László Nagy, Dávid Papp, 2018-09-18 Fifty years ago a new approach to reaction kinetics began to emerge one based on mathematical models of reaction kinetics or formal reaction kinetics Since then there has been a rapid and accelerated development in both deterministic and stochastic kinetics primarily because mathematicians studying differential equations and algebraic geometry have taken an interest in the nonlinear differential equations of kinetics which are relatively simple yet capable of depicting complex behavior such as oscillation chaos and pattern formation The development of stochastic models was triggered by the fact that novel methods made it possible to measure molecules individually Now it is high time to make the results of the last half century available to a larger audience students of chemistry chemical engineering and biochemistry not to mention applied mathematics Based on recent papers this book presents the most important concepts and results together with a wealth of solved exercises The book is accompanied by the authors Mathematica package ReactionKinetics which helps both students and scholars in their everyday work and which can be downloaded from <http://extras.springer.com> and also from the authors websites Further the large set of unsolved problems provided may serve as a springboard for individual research

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A Kinetic View of Statistical Physics Pavel L. Krapivsky, Sidney Redner, Eli Ben-Naim, 2010-11-18 Aimed at graduate students this book explores some of the core phenomena in non equilibrium statistical physics It focuses on the development and application of theoretical methods to help students develop their problem solving skills The book begins with microscopic transport processes diffusion collision driven phenomena and exclusion It then presents the kinetics of aggregation fragmentation and adsorption where the basic phenomenology and solution techniques are emphasized The following chapters cover kinetic spin systems both from a discrete and a continuum perspective the role of disorder in non equilibrium processes hysteresis from the non equilibrium perspective the kinetics of chemical reactions and the properties of complex networks The book contains 200 exercises to test students understanding of the subject A link to a website hosted by the authors containing supplementary material including solutions to some of the exercises can be found at www.cambridge.org/9780521851039 *Survival Guide to Organic Chemistry* Patrick E. McMahon, Bohdan B. Khomtchouk, Claes Wahlestedt, 2016-12-19 The Survival Guide to Organic Chemistry Bridging the Gap from General Chemistry enables organic chemistry students to bridge the gap between general chemistry and organic chemistry It makes sense of the myriad of in depth concepts of organic chemistry without overwhelming them in the necessary detail often given in a complete organic chemistry text Here the topics covered span the entire standard organic chemistry curriculum The authors describe subjects which require further explanation offer alternate viewpoints for understanding and provide hands on practical problems and

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