

Preparation of solution

0.1 N standard oxalic acid ($\text{H}_2\text{C}_2\text{O}_4 \cdot 2\text{H}_2\text{O}$) solution (100 ml)

Calculate the amount required to prepare 100 ml of 0.1 N $\text{H}_2\text{C}_2\text{O}_4 \cdot 2\text{H}_2\text{O}$: (MM: 126 g mol^{-1}).

Observations

Weight of empty weight bottle (w_1) = 2.526 g

Weight of weighing bottle + oxalic acid (w_2) = 3.162 g

Weight of weighing bottle after transferring oxalic acid (w_3) = 2.529 g

$\therefore$ amount of oxalic acid transferred ($w_2 - w_3$) = 0.633 g

Hence, exact normality of standard oxalic acid

$$= \frac{(w_2 - w_3) \times 1000}{\text{Eq. wt.} \times \text{Vol}} = 0.10 \text{ N}$$

Practical Reports On Conductometric Titrations

SA Adler



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