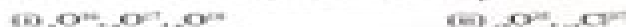


MODERN PHYSICS

1. NUCLEUS

1.1 Isotopes

The atoms of an element, which have the same atomic number but different mass numbers, are called isotopes.



1.2 Isotones

The atoms whose nuclei have same number of neutrons are called isotones.

1.3 Isobars

The atoms, which have same mass number but different atomic numbers, are called isobars.



1.4 Atomic mass unit

The atomic mass unit (a.m.u) is a very small unit of mass and it is found to be very convenient in nuclear physics.

Atomic mass unit is defined as $1/12^{\text{th}}$ of the mass of one ${}_{12}C^{12}$ atom.

According to Avogadro's hypothesis, number of atoms in 12 g of ${}_{12}C^{12}$ is equal to Avogadro number i.e. 6.023×10^{23} .

Therefore, the mass of one carbon atom (${}_{12}C^{12}$) is $\frac{12}{6.023 \times 10^{23}}$ i.e. 1.992678×10^{-26} kg.

Therefore, $1 \text{ a.m.u} = \frac{1}{12} \times 1.992678 \times 10^{-26} \text{ kg}$

or $1 \text{ a.m.u} = 1.660565 \times 10^{-27} \text{ kg}$

1.5 Energy equivalent of atomic mass unit

According to Einstein's mass-energy equivalence relation, the energy equivalent of mass m is given by $E = mc^2$

Where c is speed of light,

Suppose that $m = 1 \text{ a.m.u} = 1.660565 \times 10^{-27} \text{ kg}$

Since, $c = 2.998 \times 10^8 \text{ ms}^{-1}$, the energy equivalent of 1 a.m.u is given by $1 \text{ a.m.u} = (1.660565 \times 10^{-27} \text{ kg}) \times (2.998 \times 10^8 \text{ ms}^{-1})^2 = 1.4925 \times 10^{-10} \text{ J}$

Since, $1 \text{ MeV} = 1.602 \times 10^{-13} \text{ J}$, we have $\frac{1.4925 \times 10^{-10}}{1.602 \times 10^{-13}} \text{ eV}$

Or $1 \text{ a.m.u} = 931.5 \text{ MeV}$

1.6 Nuclear size

The volume of the nucleus is directly proportional to the number of nucleons (mass number) constituting the nucleus. If R is the radius of the nucleus having mass number A , then

$$\frac{4}{3} \pi R^3 \propto A$$

Or $R \propto A^{1/3}$ Or $R = R_0 A^{1/3}$

1.7 Nuclear density

Mass of the nucleus of the atom of mass number $A = A \text{ a.m.u}$
 $= A \times 1.660565 \times 10^{-27} \text{ kg}$. If R is radius of the nucleus, then

$$\text{Volume of nucleus} = \frac{4}{3} \pi R^3 = \frac{4}{3} \pi (R_0 A^{1/3})^3 = \frac{4}{3} \pi R_0^3 A$$

Taking $R_0 = 1.1 \times 10^{-12} \text{ m}$, we have

Density of the nucleus, $\rho = \frac{\text{mass of nucleus}}{\text{volume of nucleus}}$

$$= \frac{A \times 1.660565 \times 10^{-27}}{\frac{4}{3} \pi (1.1 \times 10^{-12})^3 \times A} \\ = 2.97 \times 10^{17} \text{ kg m}^{-3} \text{ (independent of } A)$$

Discussion :

- The density of the nuclei of all the atoms is same as it is independent of mass number.
- The high density of the nucleus ($\approx 10^{17} \text{ kg m}^{-3}$) suggests the compactness of the nucleus. Such examples of high densities are met in the form of neutron stars.

1.8 Mass defect

The difference between the sum of the masses of the nucleons constituting a nucleus and the rest mass of the nucleus is known as mass defect. It is denoted by Δm .

Let us calculate the mass defect in case of the nucleus of an atom ${}_Z X^A$. The nucleus of the atom contains Z protons and $(A-Z)$ neutrons. Therefore, if $m_n({}_Z X^A)$ is mass of the nucleus of the atom ${}_Z X^A$, then the mass defect is given by

$$\Delta m = [Zm_p + (A-Z)m_n - m_n({}_Z X^A)]$$

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