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Answers to "Why?" in the text

Page 31:

Oxygen has an atomic mass of 16 whereas it is 14 for nitrogen. The O₂ molecule is therefore heavier than the N₂ molecule. Thus, from $\frac{1}{2}m\overline{v^2} = 3(\frac{1}{2}kT)$, the rms velocity of O₂ molecules is smaller than that of N₂ molecules.

Page 34, footnote 11

For small extensions, the difference between the engineering and instantaneous strains due to a temperature change are the same. Historically, mechanical and civil engineers measured extension by monitoring the change in length, ΔL ; and the instantaneous length L was not measured. It is not trivial to measure both the instantaneous length and the extension simultaneously. However, since we know L_v and measure ΔL , the instantons length $L = L_v + \Delta L$. Is the difference important? Consider a sample of length L_v that extends to a final length L due to a temperature change from T_v to T. Let $\varepsilon = (L - L_v) / L_v = \Delta L / L_v$ be the engineering strain.

The engineering definition of strain and hence the thermal expansion coefficient is

Engineering strain =
$$\frac{\delta L}{L_a} = \lambda \delta T$$

so that thermal expansion from T_o to T gives,

$$\int_{L}^{L} \frac{dL}{L_{o}} = \int_{T}^{T} \lambda dT \qquad \therefore \qquad \frac{\Delta L}{L_{o}} = \lambda (T - T_{o}) \qquad \therefore \qquad \varepsilon = \lambda (T - T_{o}) \qquad (1)$$

where $\varepsilon = \Delta L/L_o$ is the engineering strain as defined above

Physics definition of strain and hence the thermal expansion coefficient is

Instantaneous train =
$$\frac{\delta L}{L} = \lambda \delta T$$

so that thermal expansion from To to T gives,

$$\int_{L}^{L} \frac{dL}{L} = \int_{T_{c}}^{T} \lambda dT \qquad \therefore \qquad \ln\left(\frac{L}{L_{o}}\right) = \lambda(T - T_{o}) \qquad \therefore \qquad \ln\left(1 + \varepsilon\right) = \lambda(T - T_{o}) \qquad (2)$$

We can expand the $ln(1 + \epsilon)$ term for small ϵ , so that Equation (2) essentially becomes Equation (1)

1.1 Virial theorem The Li atom has a nucleus with a +3e positive charge, which is surrounded by a full 1s shell with two electrons, and a single valence electron in the outer 2s subshell. The atomic radius of the

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Behzad Razavi

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Table of Contents Physics Solutions Manual Chapter 1

- 1. Understanding the eBook Physics Solutions Manual Chapter 1
 - The Rise of Digital Reading Physics Solutions Manual Chapter 1
 - Advantages of eBooks Over Traditional Books
- 2. Identifying Physics Solutions Manual Chapter 1
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - o Features to Look for in an Physics Solutions Manual Chapter 1
 - User-Friendly Interface
- 4. Exploring eBook Recommendations from Physics Solutions Manual Chapter 1
 - Personalized Recommendations
 - Physics Solutions Manual Chapter 1 User Reviews and Ratings
 - Physics Solutions Manual Chapter 1 and Bestseller Lists
- 5. Accessing Physics Solutions Manual Chapter 1 Free and Paid eBooks
 - Physics Solutions Manual Chapter 1 Public Domain eBooks
 - Physics Solutions Manual Chapter 1 eBook Subscription Services
 - Physics Solutions Manual Chapter 1 Budget-Friendly Options

- 6. Navigating Physics Solutions Manual Chapter 1 eBook Formats
 - o ePub, PDF, MOBI, and More
 - Physics Solutions Manual Chapter 1 Compatibility with Devices
 - Physics Solutions Manual Chapter 1 Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - o Adjustable Fonts and Text Sizes of Physics Solutions Manual Chapter 1
 - Highlighting and Note-Taking Physics Solutions Manual Chapter 1
 - Interactive Elements Physics Solutions Manual Chapter 1
- 8. Staying Engaged with Physics Solutions Manual Chapter 1
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Physics Solutions Manual Chapter 1
- 9. Balancing eBooks and Physical Books Physics Solutions Manual Chapter 1
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Physics Solutions Manual Chapter 1
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Physics Solutions Manual Chapter 1
 - Setting Reading Goals Physics Solutions Manual Chapter 1
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Physics Solutions Manual Chapter 1
 - Fact-Checking eBook Content of Physics Solutions Manual Chapter 1
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements

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