



NATIONAL
EXAM CERTIFICATE

GRADE 11

NOVEMBER 2020

PHYSICAL SCIENCES P2
(CHEMISTRY)
EXEMPLAR

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QUESTION 3

PHYSICAL SCIENCES PAPER 2

INTERMOLECULAR FORCES

GRADE 11

NOVEMBER 2020

QUESTION NO. _____ INITIAL ANSWERS _____

QUESTION 3 starts on PAPER 2 page 2.

3.1 The boiling points in the table below were obtained during an investigation into the boiling points of pure 1-hexanol. The table below lists the boiling points of pure 1-hexanol and pure 1-hexanol with 10% water.

Compound	Boiling point (°C)	Boiling point (°C)
A	158	158
B	158	158

- 3.1.1 Define the term boiling point. (2)
- 3.1.2 In which column are compounds A and B in (1) and (2) lists of boiling points? (2)
- 3.1.3 Name the type of intermolecular force that water shows between molecules of pure compounds A and B due to the polar nature of their molecules. (2)
- 3.1.4 What (10%) of the compound B in (2) was 1-hexanol? Explain your answer. (2)
- 3.1.5 What compound B in (2) did not have hydrogen bonds? (2)
- 3.1.6 Explain how the boiling point of compound B in (2) compares to that of compound A and B due to the nature of the intermolecular forces between their molecules. (2)

3.2 A 100 mL sample of water in the table has a smaller volume than that of compound A and B due to the nature of the intermolecular forces between their molecules. (2)

- 3.2.1 Explain why the boiling point of compound B in (2) is higher than that of compound A and B due to the nature of the intermolecular forces between their molecules. (2)

3.3 Additive (1) is added to each column of pure compounds (2) and the boiling points are lower than those of pure 1-hexanol. (2)



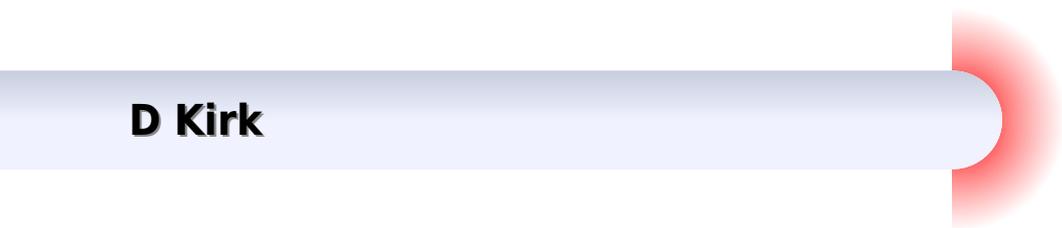
- 3.3.2 In which beaker (1) or (2), will the water boil? (2)
- 3.3.3 Explain the reason for (3.3.2) based on the nature of the intermolecular forces between their molecules. (2)

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