Name:

 Circle the element with the largest atomic radius and put a square around the element with the smallest atomic radius:

Cu K Ni Br

- a. Explain why you made these choices: All of the elements are in the same period. The trend in atomic radius as you go across a period is DECREASING. Therefore, the element on the far left (K) is the largest, and the element on the far right (Br) is the smallest.
- Circle the element with the highest ionization energy and put a square around the element with the lowest ionization energy:

Cu K Ni (Br)

- a. Explain why you made these choices: All of the elements are in the same period. The trend in ionization energy as you go across a period is INCREASING. Therefore, the element on the far left (K) has the lowest ionization energy, and the element on the far right (Br) has the highest ionization energy.
- Circle the element with the highest electronegativity and put a square around the element with the lowest electronegativity:

Cu K Ni Br

- a. Explain why you made these choices: All of the elements are in the same period. The trend in electronegativity as you go across a period is INCREASING. Therefore, the element on the far left (K) has the lowest electronegativity, and the element on the far right (Br) has the highest electronegativity.
- 4. For each of the following groups: Circle the element with the largest atomic radius and put a square around the element with the smallest atomic radius:

5.

- a. O C (Bc) Nc Same Period
  b. Na Rb (Fr) H Same Group
  c. (Pb) C Sn Si Same Group
  d. Au W S (Fr) Nc Zn Challenge
- For each of the following groups: Circle the element with the highest ionization energy and put a square around the element with the lowest ionization energy:
  - a. O C Bc No Same Period
    b. Na Rb Fr H Same Group
    c. Pb C Sn Si Same Group
    d. Au W S Fr No Zn Challenge

# **Periodic Trends Worksheet Answers**

**George Graybill** 

#### **Periodic Trends Worksheet Answers:**

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#### Table of Contents Periodic Trends Worksheet Answers

- 1. Understanding the eBook Periodic Trends Worksheet Answers
  - The Rise of Digital Reading Periodic Trends Worksheet Answers
  - Advantages of eBooks Over Traditional Books
- 2. Identifying Periodic Trends Worksheet Answers
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Periodic Trends Worksheet Answers
  - User-Friendly Interface
- 4. Exploring eBook Recommendations from Periodic Trends Worksheet Answers
  - Personalized Recommendations
  - Periodic Trends Worksheet Answers User Reviews and Ratings
  - Periodic Trends Worksheet Answers and Bestseller Lists
- 5. Accessing Periodic Trends Worksheet Answers Free and Paid eBooks
  - Periodic Trends Worksheet Answers Public Domain eBooks
  - Periodic Trends Worksheet Answers eBook Subscription Services
  - Periodic Trends Worksheet Answers Budget-Friendly Options

- 6. Navigating Periodic Trends Worksheet Answers eBook Formats
  - o ePub, PDF, MOBI, and More
  - Periodic Trends Worksheet Answers Compatibility with Devices
  - Periodic Trends Worksheet Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - o Adjustable Fonts and Text Sizes of Periodic Trends Worksheet Answers
  - Highlighting and Note-Taking Periodic Trends Worksheet Answers
  - Interactive Elements Periodic Trends Worksheet Answers
- 8. Staying Engaged with Periodic Trends Worksheet Answers
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Periodic Trends Worksheet Answers
- 9. Balancing eBooks and Physical Books Periodic Trends Worksheet Answers
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Periodic Trends Worksheet Answers
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Periodic Trends Worksheet Answers
  - Setting Reading Goals Periodic Trends Worksheet Answers
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Periodic Trends Worksheet Answers
  - Fact-Checking eBook Content of Periodic Trends Worksheet Answers
  - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
- 14. Embracing eBook Trends
  - Integration of Multimedia Elements

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