

PERCENTAGE COMPOSITION

Name _____

Determine the percentage composition of each of the compounds below:

1. Al_2SiO_5
 $\text{Al} = \frac{2 \times 27}{178} \times 100 = 30.34\%$
 $\text{Si} = \frac{28}{178} \times 100 = 15.73\%$
 $\text{O} = \frac{5 \times 16}{178} \times 100 = 53.93\%$

2. $\text{Fe}_2(\text{SO}_4)_3$
 $\text{Fe} = \frac{2 \times 56}{504} \times 100 = 22.22\%$
 $\text{S} = \frac{3 \times 32}{504} \times 100 = 18.89\%$
 $\text{O} = \frac{12 \times 16}{504} \times 100 = 58.89\%$

3. H_2SO_4
 $\text{H} = \frac{2 \times 1}{98} \times 100 = 2.04\%$
 $\text{S} = \frac{32}{98} \times 100 = 32.65\%$
 $\text{O} = \frac{4 \times 16}{98} \times 100 = 65.31\%$

4. Mg_2SiO_4
 $\text{Mg} = \frac{2 \times 24}{160} \times 100 = 30\%$
 $\text{Si} = \frac{28}{160} \times 100 = 17.5\%$
 $\text{O} = \frac{4 \times 16}{160} \times 100 = 52.5\%$

5. $\text{Ba}_3(\text{PO}_4)_2$
 $\text{Ba} = \frac{3 \times 137}{601} \times 100 = 68.73\%$
 $\text{P} = \frac{2 \times 31}{601} \times 100 = 10.32\%$
 $\text{O} = \frac{8 \times 16}{601} \times 100 = 20.95\%$

6. Al_2SO_4
 $\text{Al} = \frac{2 \times 27}{178} \times 100 = 30.34\%$
 $\text{S} = \frac{32}{178} \times 100 = 18.03\%$
 $\text{O} = \frac{4 \times 16}{178} \times 100 = 51.63\%$

Solve the following problems:

- How much oxygen is evolved when 200 g of H_2O_2 is decomposed?
 $\text{H}_2\text{O}_2 \rightarrow \text{H}_2\text{O} + \text{O}_2$
 $2 \times 34 = 2 \times 18 + 2 \times 16$
 $68 = 36 + 32$
 $\text{O}_2 = \frac{32}{68} \times 200 = 94.12 \text{ g}$
- How much SO_2 can be recovered from 200 g of FeS_2 ?
 $\text{FeS}_2 \rightarrow \text{Fe} + 2\text{SO}_2$
 $120 = 56 + 2 \times 64$
 $\text{SO}_2 = \frac{128}{120} \times 200 = 213.33 \text{ g}$
- How much H_2SO_4 can be produced from 100 g of Fe_2O_3 ?
 $\text{Fe}_2\text{O}_3 + 3\text{H}_2\text{SO}_4 \rightarrow \text{Fe}_2(\text{SO}_4)_3 + 3\text{H}_2\text{O}$
 $160 + 3 \times 98 = 344 + 54$
 $\text{H}_2\text{SO}_4 = \frac{294}{160} \times 100 = 183.75 \text{ g}$

Percentage Composition 54 Chemistry If8766 Answers

Jicheng Xie



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