

May 2009 TZ1 Question 1

- (b) For part of her extended essay investigation into the efficiency of the process, a student reacted a pure sample of a vegetable oil (where $R = C_{17}H_{33}$) with methanol. The raw data recorded for the reaction is below.

Mass of oil	= 1013.0 g
Mass of methanol	= 200.0 g
Mass of sodium hydroxide	= 3.5 g
Mass of biodiesel produced	= 811.0 g

The relative molecular mass of the oil used by the student is 885.6. Calculate the amount (in moles) of the oil and the methanol used, and hence the amount (in moles) of excess methanol.

[3]

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May 2009 TZ2 Question 1

- (b) A student reacted some salicylic acid with excess ethanoic anhydride. Impure solid aspirin was obtained by filtering the reaction mixture. Pure aspirin was obtained by recrystallization. The following table shows the data recorded by the student.

Mass of salicylic acid used	3.15 ± 0.02 g
Mass of pure aspirin obtained	2.50 ± 0.02 g

- (i) Determine the amount, in mol, of salicylic acid, $C_6H_4(OH)COOH$, used.

[2]

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- (ii) Calculate the theoretical yield, in g, of aspirin, $C_6H_4(OCOCH_3)COOH$.

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9 the sketch shown below is for a commercial sign it was drawn to 1 10 of its true size what is the area of the actual sign prove your answer by showing all calculations note see more

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web activity 5 1 calculating properties of shapes intro finding the surface area of shapes to know how much stuff you need to put on the shape without putting too much or little