



Lab Data



*Verify your calculation. In your set-up, do your units cancel?
Did you report your data to the correct number of significant figures?*

Moles of magnesium (mol) 0.00062

Temperature of water (°C) 22.0

Temperature of water (K) 295.15

Vapor pressure of water (mmHg) 19.8

Barometric pressure (mmHg) 770.0

Observations

There is bubble formation

Volume of hydrogen gas collected (mL) 15.00

Volume of hydrogen gas collected (L)

Pressure of hydrogen gas (mmHg)

Pressure of hydrogen gas (atm)

Moles of hydrogen gas produced (mol) 0.00062

Ideal gas constant ($\text{L}\cdot\text{atm}\cdot\text{mol}^{-1}\cdot\text{K}^{-1}$)



[How to calculate ideal gas constant](#)



MY

Ideal Gas Constant Lab 38 Answers

**Eugene LeMay, Jr., Herbert Beall, Karen
M. Robblee, Douglas C.
Brower, Prentice-Hall Staff**

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