

1-6, 8, 9, 11

Hydrolysis of Salts and pH of Buffer Solutions

Pre-lab Questions

Before beginning this experiment in the laboratory, you should be able to answer the following questions.

1. Define Brønsted-Lowry acids and bases.

According to Brønsted-Lowry definitions, an acid is a proton donor and a base is a proton acceptor.

2.

- Which of the following ions will react with water in a hydrolysis reaction:

K^+ , Ba^{2+} , Cu^{2+} , Zn^{2+} , F^- , SO_3^{2-} , and Cl^- ?

Cu^{2+} , Zn^{2+} , F^- and SO_3^{2-} will undergo hydrolysis.

3.

- For those ions in question 2 that undergo hydrolysis, write net ionic equations for the hydrolysis reaction.

Example: $Cu^{2+}(aq) + 2H_2O(l) \rightleftharpoons Cu(OH)_2(s) + 2H^+(aq)$.

4.

- The K_a for HCN is 4.9×10^{-10} . What is the value of K_b for CN^- ?

$$K_b = \frac{1.0 \times 10^{-14}}{4.9 \times 10^{-10}} = 2.0 \times 10^{-5}$$

5. What are the conjugate base and conjugate acid of $H_2PO_4^-$?

The conjugate base is HPO_4^{2-} and the conjugate acid is H_3PO_4 .

6.

- From what acid and what base were the following salts made: K_2SO_4 , NH_4Cl , and $Ca(NO_3)_2$?

For $CaSO_4$, may be made from $Ca(OH)_2$ and H_2SO_4 .

7.

- Define the term *salt*.

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Daniel Hernandez-Patlan, Bruno Solis-Cruz, Guillermo Tellez-Isaias

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Green Chemistry Laboratory Manual for General Chemistry Sally A. Henrie, 2015-03-18 Green chemistry involves designing novel ways to create and synthesize products and implement processes that will eliminate or greatly reduce negative environmental impacts Providing educational laboratory materials that challenge students with the customary topics found in a general chemistry laboratory manual this lab manual enables students to see how green chemistry principles can be applied to real world issues Following a consistent format each lab experiment includes objectives prelab questions and detailed step by step procedures for performing the experiments Additional questions encourage further research about how green chemistry principles compare with traditional more hazardous experimental methods **40 Inquiry Exercises for**

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