# CHAPTER 18: ELECTRIC CHARGE AND ELECTRIC FIELD

#### 18.1 STATIC ELECTRICITY AND CHARGE: CONSERVATION OF CHARGE

- Common static electricity involves charges ranging from nanocoulombs to microcoulombs. (a) How many electrons are needed to form a charge of -2.00 nC? (b) How many electrons must be removed from a neutral object to leave a net charge of 0.500 µC?
- Solution (a) Since one electron has charge of  $q_r = -1.6 \times 10^{-19} \, \text{C}$ , we can determine the number of electrons necessary to generate a total charge of -2.00 nC by using the equation  $Nq_r = Q$ , so that  $N = \frac{Q}{q_r} = \frac{-2.00 \times 10^{-9} \, \text{C}}{-1.60 \times 10^{-19} \, \text{C}} = \frac{1.25 \times 10^{10} \, \text{electrons}}{1.25 \times 10^{10} \, \text{electrons}}$ .
  - (b) Similarly we can determine the number of electrons removed from a neutral object to leave a charge of 0.500 µC of charge so that:

$$N = \frac{-0.500 \times 10^{-6} \text{ C}}{-1.60 \times 10^{-19} \text{ C}} = 3.13 \times 10^{12} \text{ electrons}$$

#### 18.2 CONDUCTORS AND INSULATORS

- A 50.0 g ball of copper has a net charge of 2.00 μC. What fraction of the copper's electrons has been removed? (Each copper atom has 29 protons, and copper has an atomic mass of 63.5.)
- Solution Recall that Avogadro's number is  $N_x = 6.02 \times 10^{23}$  atoms/mole. Now we need to determine the number of moles of copper that are present. We do this using the mass and the atomic mass:  $n = \frac{m}{A} = \frac{50.0 \text{ g}}{63.5 \text{ g/mol}}$

So since there are 29 protons per atom we can determine the number of protons,

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**Atam Parkash Arya** 

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