

DEVOIR DE CONTROLE N°2

Date: 10 -02-2020

Durée : 2H

- le sujet comporte 2 exercices de chimie et 2 exercices de physique.
- On donnera l'expression littérale avant de passer à l'application numérique.
- L'utilisation de la calculatrice scientifique est autorisée.



CHIMIE : (9 POINTS)

Exercice n°1 :

Un acide carboxylique (A) de masse molaire $M = 74 \text{ g.mol}^{-1}$.

- Déterminer la formule semi développée et le nom de (A).
- On dissout une masse $m=0,74\text{g}$ de (A) dans l'eau distillée afin de préparer 100mL d'une solution (S) d'acide de concentration molaire C. Le pH de cette solution est égal à 3.
 - Calculer la concentration molaire C.
 - Comparer C et $[H_3O^+]$. Ecrire son équation de dissociation dans l'eau.
 - Déduire la molarité des espèces chimiques présentes dans la solution (sauf l'eau).
On rappelle qu'à 25°C : $[H_3O^+][OH^-]=10^{-14}$
- On fait réagir un volume $V^*=50\text{mL}$ de la solution (S) avec $m^*=0,1308\text{g}$ de zinc, il se dégage un gaz qui provoque une légère détonation en présence d'une flamme.
Quel est le gaz dégagé ? Calculer son volume.
- On fait réagir 5 moles de (A) avec 5 moles d'alcool (B) qui donne à la fin de la réaction 3 moles d'un produit (C) formé par 6 atomes de carbone de plus 3 moles d'eau.
 - Déterminer la formule brute de l'alcool (B) ainsi que la formule semi-développée de (C) et son nom.
 - Dégager de ce qui précède la propriété de cette réaction.

0,75B

0,75B

0,75B

0,75B

1C

1C

0,5B

Exercice n°2 :

La combustion d'une masse $m_1=2\text{g}$ d'un alcool (A) $C_nH_{2n+2}O$ fournit $m_2=5\text{g}$ de dioxyde de carbone.

- Ecrire l'équation générale de la réaction de combustion.
- Montrer que la masse molaire de l'alcool (A) peut s'écrire $M=17,6n$.
- Déduire alors la formule brute de (A)
- L'oxydation ménagée de (A) par le bichromate de potassium ($2K^+$, $Cr_2O_7^{2-}$) en milieu acide donne un composé (B) qui réagit avec le 2,4 DNPH mais ne réagit pas avec le réactif de Schiff.
 - Préciser la fonction chimique de (B) ?
 - Déduire alors tous les isomères possibles de (A).
- Ecrire l'équation de la réaction d'oxydation ménagée en utilisant les formules brutes.
- La déshydratation intramoléculaire de (A) donne un seul composé organique (C). Identifier les composés organiques (A), (B) et (C).

0,5A

0,5 C

0,5B

0,25A

0,75A

1C

1C

On donne g. mol^{-1} : $M(C)=12$; $M(O)=16$; $M(H)=1$; $M(Zn)=65,4$ et $V_m=24 \text{ L.mol}^{-1}$



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