

## Determination of Iron in Ore by Redox Titration

Oxidation-reduction reactions are suitable for titration reactions if they are rapid and quantitative. The end point can be seen by indicators or by instrumental methods such as potentiometry.

### Potentiometric Titration of Ferrous Ion

The reaction of ferrous ion with ceric ion meets the above requirements.



The course of the titration may be followed potentiometrically with a platinum electrode. The following cell may be used:



The potential of the Pt electrode must obey the Nernst equation since the half reactions between  $\text{Fe}^{3+}$  and  $\text{Fe}^{2+}$  and between  $\text{Ce}^{4+}$  and  $\text{Ce}^{3+}$  are nearly reversible.

$$E_{\text{Pt}} = E^{\circ}(\text{Fe}^{3+}, \text{Fe}^{2+}) + 0.05916 \log \left( \frac{[\text{Fe}^{3+}]}{[\text{Fe}^{2+}]} \right) = E^{\circ}(\text{Ce}^{4+}, \text{Ce}^{3+}) + 0.05916 \log \left( \frac{[\text{Ce}^{4+}]}{[\text{Ce}^{3+}]} \right) \quad (2)$$

Equation 2 states that the Pt electrode can have only one potential, and the ratios of  $[\text{Fe}^{3+}]/[\text{Fe}^{2+}]$  and  $[\text{Ce}^{4+}]/[\text{Ce}^{3+}]$  must be such that the Nernst equation is satisfied for both half reactions. Reaction 1 will occur spontaneously to such an extent that equation 2 will be true. The ratios  $[\text{Fe}^{3+}]/[\text{Fe}^{2+}]$  and  $[\text{Ce}^{4+}]/[\text{Ce}^{3+}]$  change greatly at the equivalence point. The equivalence point is determined from the break in the titration curve.

### Indicator Titration of Ferrous Ion

Specific indicators and true oxidation-reduction indicators are available for redox titrations. Specific indicators are substances that react with one or more of the participants in a titration to produce a color change. A classic example is starch that forms a dark blue complex with  $\text{I}_2$ . In redox titrations with  $\text{I}_2$  as the titrant, blue color develops to mark the end point of the reaction when  $\text{I}_2$  begins to accumulate in the solution. Thiocyanate is an indicator that forms a red complex with  $\text{Fe}^{3+}$  but not with  $\text{Fe}^{2+}$ , and thus can be used to indicate the formation or disappearance of  $\text{Fe}^{3+}$ . Sometimes a redox reagent itself is deeply colored and can be used to indicate the end point. For example,  $\text{MnO}_4^-$  is a purple ion often used as an oxidizing agent. The first trace of purple color indicates that the  $\text{MnO}_4^-$  is no longer being reduced by the titration reaction.

A true oxidation-reduction indicator is a substance that has different colors when in different oxidation states.



# Determination Of Iron In Ore By Redox Titration Chemistry

**Wolfgang Guggemos**



## **Determination Of Iron In Ore By Redox Titration Chemistry:**

Analytical Chemistry of Complex Matrices ,2013-03-08      Analytical Chemistry Gary D. Christian,Purnendu K. Dasgupta,Kevin A. Schug,2013-10-07 With the 7th Edition of Analytical Chemistry renowned chemists Purnendu Sandy Dasgupta and Kevin Schug both of the University of Texas Arlington join the author team The new edition focuses on more in depth coverage of the principles and techniques of quantitative analysis and instrumental analysis aka Analytical Chemistry The goal of the text is to provide a foundation of the analytical process tools and computational methods and resources and to illustrate with problems that bring realism to the practice and importance of analytical chemistry It is designed for undergraduate college students majoring in chemistry and in fields related to chemistry ,      **Comprehensive Engineering Chemistry** Devender Singh,Satish Kumar Vats,2013-12-30 This book is designed to meet the requirement of the students of B Tech and B E students The book discusses in detail the following topics Thermodynamics Phase Rule Water and its Treatment Corrosion and its Prevention Lubrication and Lubricants Polymer and Polymerization and Analytical Methods The book is suitably illustrated with diagrams and a number of solved numerical examples from different universities are included to make the text more exhaustive and understandable Practical part is also appended at the end of the book      **Engineering Chemistry with Laboratory Experiments** Kaurav M. S.,2011      **Chemical Equilibrium and Analysis** Richard W. Ramette,1981      **Quantitative Chemical Analysis** Na Li,John J Hefferren,Ke'an Li,2013-04-26 This book covers both fundamental and practical aspects of chemical analysis Data Process and Analysis Chemical Equilibria and Volumetric titrations Gravimetry Spectrophotometry Sample Preparation and Separation Methods in Quantitative Analysis It was written with the rich tradition of teaching at Peking University College of Chemistry and edited by an American professor who was personally sensitive to the needs of students learning science from traditional chemistry textbooks written in English Many examples and illustrative problems in this text have been taken from previous textbooks by the Peking University Team Teaching Program The book can be used as a starter in analytical chemistry which is fundamental and the base upon which chemistry is built Traditional chapters of initial learning in analytical chemistry are included such as volumetric gravimetric and separation methods the book also includes key chapters on problem solving relating to recent progress in analytical chemistry      *R - Sm, Berichtsjahre 1932 bis 1962, Lieferung 5* Johann Christian Poggendorff,2022-09-06 Keine ausf hrliche Beschreibung f r R Sm Berichtsjahre 1932 bis 1962 Lieferung 5 verf gbar      **SME Mineral Processing and Extractive Metallurgy Handbook** Courtney A. Young,2019-02-01 This landmark publication distills the body of knowledge that characterizes mineral processing and extractive metallurgy as disciplinary fields It will inspire and inform current and future generations of minerals and metallurgy professionals Mineral processing and extractive metallurgy are atypical disciplines requiring a combination of knowledge experience and art Investing in this trove of valuable information is a must for all those involved in the industry students engineers mill managers and operators

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**A Manual for the Chemical Analysis of Metals**, **A Level Salters Advanced Chemistry for OCR B** University of York, 2016-05-05

Written by the University of York project team for Salters Advanced Chemistry, this Student Book supports and extends students through the new linear course while delivering the breadth, depth and skills needed to succeed in the new A Level and beyond. It develops true subject knowledge while also developing essential exam skills. The fourth edition combines the Chemical Storyline and Chemical Ideas into a single integrated volume for the first time, providing ideal support for the new specification.

Engineering Chemistry Laboratory Manual Dr Manoj Kumar Solanki, 2019-03-20

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**Analytical Chemistry, International Adaptation** Gary D. Christian, Purnendu K. Dasgupta, Kevin A. Schug, 2020

With the 7th Edition of Analytical Chemistry, renowned chemists Purnendu K. Dasgupta and Kevin Schug, both of the University of Texas at Arlington, join the author team. The new edition focuses on more in-depth coverage of the principles and techniques of quantitative analysis and instrumental analysis, aka Analytical Chemistry. The goal of the text is to provide a foundation of the analytical process, tools and computational methods and resources and to illustrate with problems that bring realism to the practice and importance of analytical chemistry. It is designed for undergraduate college students majoring in chemistry and in fields related to chemistry.

Quantitative Chemical Analysis Leicester Forsyth Hamilton, 1958

A Brief Introduction to Quantitative Chemical Analysis Robert Blanchard Fischer, Dennis G. Peters, 1969

*Nature Of Chemistry Volume - 2* Aditya Tripathi, **AP Chemistry with Online Tests** Neil D. Jespersen, Pamela Kerrigan, 2020-07-07

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