GASEOUS STATE

Measurable properties of gases :

1. **Mass :**
   Def. The gases do possess mass. The mass of gas is generally used in the form of number of moles which is related as :

2. **Volume :**
   Def. Volume of gas is nothing but volume of the container in which it is present. Relation between different units of volume
   $1 \, m^3 = 10^3 \, dm^3 = 10^3 \, litre = 10^6 \, cm^3 = 10^6 \, ml = 10^9 \, mm^3$.

3. **Temperature :**
   Def. Degree of hotness or coldness of a body is measured by temperature

   $$\frac{C}{100} = \frac{K - 273}{100} = \frac{F - 32}{180}$$

   C – Celcius scale,   K – Kelvin scale,   F – Fahrenheit scale

   Note : In all the problems of gaseous state (i.e. in all gas law equations), temperature must be expressed in kelvin scale. i.e., **t °C + 273 = TK**

4. **Pressure :**
   Def. Force acting per unit area

   $$P = \frac{F}{A}$$

   Units :
   CGS   : $dyne/cm^2$
   MKS   : $Newton/m^2$ $(1N/m^2 = 1Pa)$
   Relation   : $1 \, N/m^2 = 10 \, dyne/cm^2$

   Units of pressure :
   1 atm   = 76 cm of Hg
           = 760 mm of Hg
           = 760 torr
           = $1.01325 \times 105 \, N/m^2$
           = 101.325 kPa
           = 1.01325 bar
           = 14.7 lb/In² (Psi)
           = 10.33 meters of H_2O

DENSITY OF GASES

Absolute density
(mass per unit volume)

$$d = \frac{m}{v}$$

Relative density
(Relative to hydrogen turned as vapour density)

$$VD = \frac{molecular \, mass}{2}$$

Boyle's law and measurement of pressure :

Statement :
For a fixed amount of gas at constant temperature, the volume occupied by the gas is inversely proportional to the pressure applied on the gas or pressure of the gas.

$$V \, \alpha \, \frac{1}{P}$$

hence PV = constant
this constant will be dependent on the amount of the gas and temperature of the gas.

$$P_1 V_1 = P_2 V_2$$

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