

INV. 5.1.1 MOLAR ENTHALPY of SOLUTION - MEDICAL COLD PACKS Part 2

- Describe the difference between heat and temperature.

Heat energy is the energy due the position and random motion of particles. It is thermal energy being transferred. Temperature is a measure of average kinetic energy.

- For the compounds used in this experiment (Nelson Chemistry 12 p. 347):

(a) Calculate the number of moles present in a 10.00 g sample.

$$m_{\text{NH}_4\text{Cl}} = 10.0 \text{ g} \quad M_{\text{NH}_4\text{Cl}} = 53.50 \text{ g/mol} \quad H_s(\text{NH}_4\text{Cl}) = 0.277 \text{ kJ/g}$$

$$n_{\text{NH}_4\text{Cl}} = \frac{m}{M} = 10.0 \text{ g} \times \frac{1 \text{ mol NH}_4\text{Cl}}{53.50 \text{ g}} = 0.187 \text{ mol}$$

Similarly;

$$\text{KNO}_3 = 0.0989 \text{ mol}$$

$$\text{NH}_4\text{NO}_3 = 0.125 \text{ mol}$$

$$\text{NaC}_2\text{H}_3\text{O}_2 \cdot 3 \text{ H}_2\text{O} = 0.0735 \text{ mol}$$

$$\text{KCl} = 0.134 \text{ mol}$$

(b) Calculate the molar enthalpy

$$m_{\text{NH}_4\text{Cl}} = 10.0 \text{ g} \quad M_{\text{NH}_4\text{Cl}} = 53.50 \text{ g/mol} \quad H_s(\text{NH}_4\text{Cl}) = 0.277 \text{ kJ/g}$$

$$\begin{aligned} \Delta H_{\text{sol}}(\text{NH}_4\text{Cl}) &= H_s(\text{NH}_4\text{Cl}) \times M_{\text{NH}_4\text{Cl}} \\ &= \frac{0.277 \text{ kJ}}{\text{g}} \times \frac{53.50 \text{ g}}{1 \text{ mol NH}_4\text{Cl}} = 14.8 \text{ kJ/mol} \end{aligned}$$

Similarly;

$$\text{KNO}_3 = 34.9 \text{ kJ/mol}$$

$$\text{NH}_4\text{NO}_3 = 25.7 \text{ kJ/mol}$$

$$\text{NaC}_2\text{H}_3\text{O}_2 \cdot 3 \text{ H}_2\text{O} = 19.6 \text{ kJ/mol}$$

$$\text{KCl} = 17.2 \text{ kJ/mol}$$

(c) Write the thermochemical chemical equation for your compound's dissociation in water.



- When a substance dissolves its heat content, or enthalpy, will either increase or decrease. This change in enthalpy is observed as a change in temperature of the water in the calorimeter. Sketch a graph showing the change in enthalpy for the dissolving of an ionic compound which is:

(i) exothermic



(ii) endothermic



Enthalpy Calorimetry Name Chem Worksheet 16 4

Michael Hoelscher

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