

1.36 Determine the mass of air in a  $2 \text{ m}^3$  tank if the air is at room temperature,  $20^\circ\text{C}$ , and the absolute pressure within the tank is  $200 \text{ kPa}$  (abs).

$$m = \rho V \text{ where } V = 2 \text{ m}^3 \text{ and}$$

$$\rho = p/RT \text{ with } T = 20^\circ\text{C} = (20 + 273) \text{ K} = 293 \text{ K}$$

$$\text{and } p = 200 \text{ kPa} = 200 \times 10^3 \frac{\text{N}}{\text{m}^2}$$

Thus,

$$\rho = (200 \times 10^3 \frac{\text{N}}{\text{m}^2}) / [(2.869 \times 10^2 \frac{\text{N}\cdot\text{m}}{\text{kg}\cdot\text{K}})(293 \text{ K})]$$

$$= 2.38 \frac{\text{kg}}{\text{m}^3}$$

Hence,

$$m = \rho V = 2.38 \frac{\text{kg}}{\text{m}^3} (2 \text{ m}^3) = \underline{\underline{4.76 \text{ kg}}}$$

# Fundamentals Of Fluid Mechanics 6th Edition Solution Manual Si

**Philip M. Gerhart, Richard J. Gross, John  
I. Hochstein**



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