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web where the salt is the conjugate base of the acid the four parts of the titration curve are described below and you should
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web solution a titrant volume 0 ml the solution ph is due to the acid ionization of HCl because this is a strong acid the
ionization is complete and the hydronium ion molarity is 0.100 M the ph of the solution is then $ph = -\log 0.100 = 1.000$ $ph = -\log 0.100 = 1.000$ b titrant volume 12.50 ml

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Volume of NaOH added (mL)	pH
4	6.4
6	6.5
9	5.9
5	5.9

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