

(ii)

$$15.125 \text{ mol NaOH} \cdot x \text{ mol NaOH} = \frac{15.125 \text{ mol NaOH}}{1} \times \frac{1}{1000} = 0.015125 \text{ mol NaOH}$$

$$\therefore 15.125 \text{ mol NaOH} = \frac{15.125 \text{ mol NaOH}}{1} \times \frac{1}{1000} = \frac{1.5125 \text{ g NaOH}}{1 \text{ g NaOH}} \times \frac{1 \text{ mol}}{40 \text{ g}} = \frac{1.5125 \text{ mol}}{40} = \frac{0.0378125 \text{ mol NaOH}}{1}$$

$$\therefore \text{volume of } 0.1 \text{ M NaOH} = \frac{0.0378125 \text{ mol NaOH}}{0.1 \text{ M NaOH}} = 0.378125 \text{ L}$$

Thus, 0.378125 L of the concentrated reagent will dilute to a final volume of 0.378125 L, giving

answer.

(iii)

$$10 \text{ mol NaOH} = \frac{10 \text{ mol NaOH}}{1} \times \frac{1}{1000} = 0.01 \text{ mol NaOH}$$

$$10 \text{ mol NaOH} = \frac{10 \text{ mol NaOH}}{1} \times \frac{1}{1000} = 0.01 \text{ mol NaOH}$$

$$10.4125 \text{ mol NaOH} = \frac{10.4125 \text{ mol NaOH}}{1} \times \frac{1}{1000} = 0.0104125 \text{ mol NaOH}$$

Therefore, 1.4125 g NaOH is enough to give a final volume of 0.378125 L.

Q.10

$$15.125 \text{ mol NaOH} = \frac{15.125 \text{ mol NaOH}}{1} \times \frac{1}{1000} = 0.015125 \text{ mol} = 0.015125 \times 10^{-3} \text{ mol NaOH}$$

$$15.125 \text{ mol NaOH} = \frac{15.125 \text{ mol NaOH}}{1} \times \frac{1}{1000} = 0.015125 \text{ mol} = 0.015125 \times 10^{-3} \text{ mol NaOH}$$

Therefore, each mole of $\text{NaOH} \cdot x$ requires three moles HCl , HCl requires the limiting reagent. Thus,

$$0.015125 \times 10^{-3} \text{ mol NaOH} = \frac{0.015125 \times 10^{-3} \text{ mol NaOH}}{1} \times \frac{3 \text{ mol HCl}}{1 \text{ mol NaOH}} = 0.045375 \text{ mol HCl}$$

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