

[CLASS XII CHEMISTRY PRACTICALS]

Evaluation Scheme for Examination Marks

Volumetric Analysis	08
Salt Analysis	08
Content Based Experiment	06
Project Work	04
Class record and viva	04
Total	30

- Notes:-**
1. Chemical Equations of Experiment 5 to 11 are to be written on blank pages.
 2. Observation table of experiment 12 to 18 are to be drawn on blank pages.
 3. Investigatory Project work (4 marks) is also included in the practical syllabus. For project work, contact the teacher for the topic.
 4. Project report should be hand written.
 5. Start each experiment from a new page.

EXPERIMENT – 1 **Classification of Anions**

Group	Group Reagent	Observation	Inference
A	Dilute H_2CO_3	a) Colourless, odourless gas with brisk effervescence (CO_2) which turns lime water milky. b) Colourless gas with rotten egg like smell (H_2S) which turns lead acetate paper black. c) Colourless gas with smell of burning sulphur (SO_2) which turns acidified dichromate paper green. d) Brown coloured gas (NO_2) which turns ferrous sulphate solution black or brown. e) Colourless gas with vinegar like smell.	CO_3^{2-} (Carbonate) S^{2-} (sulphide) SO_3^{2-} (Sulphite) NO_2^- (Nitrite) CH_3COO^- (Acetate)
B	Conc. H_2SO_4	a) Colourless pungent smelling gas (HCl) which gives white dense fumes with glass rod dipped in NH_4OH . b) Violet coloured vapours (I_2) which turns starch paper blue. c) Reddish brown gas (NO_2) having pungent smell (On adding copper turning, fumes becomes intense) d) Brown colour gas with pungent smell (Br_2) which turns starch paper yellow. e) Colourless, odourless gas with brisk effervescence ($CO + CO_2$) which turns lime water milky and burns on the mouth of test tube with blue flame.	Cl^- (Chloride) I^- (Iodide) NO_2^- (Nitrate) Br^- (Bromide) $C_2O_4^{2-}$ (Oxalate)
C	$BaCl_2$	White ppt. of $BaSO_4$ is formed.	SO_4^{2-} (sulphate)
D	Ammonium molybdate 3 ($NH_4)_2MoO_4$	Canary yellow ppt. of phosphor ammonium molybdate (NH_4), PO_4 1:2 MoO_4 . 6H ₂ O	PO_4^{3-} (phosphate)

EXPERIMENT – 2

Classification of Cations

Group / Reagent	Group Reagent	Radical	FT/Smell Smell of NH_4	Colour
I	$NaOH$	NH_4^+		-
II	Dil. HCl	Pb^{2+}	$PbCl_2$	White
II	H_2S gas in acidic Medium	Pb^{2+}	PbS	Black
		As^{3+}	As_2S_3	Yellow
		Cr^{3+}	Cr_2S_3	Black
		Cr^{6+}	CrS	Yellow
III	NH_4Cl (s) in presence of NH_4OH	Fe^{3+}	$Fe(OH)_3$	Light green
		Fe^{2+}	$Fe(OH)_2$	Reddish brown
IV	H_2S gas in basic medium	Al^{3+}	$Al(OH)_3$	Colourless white
		Sn^{2+}	SnS	Black

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