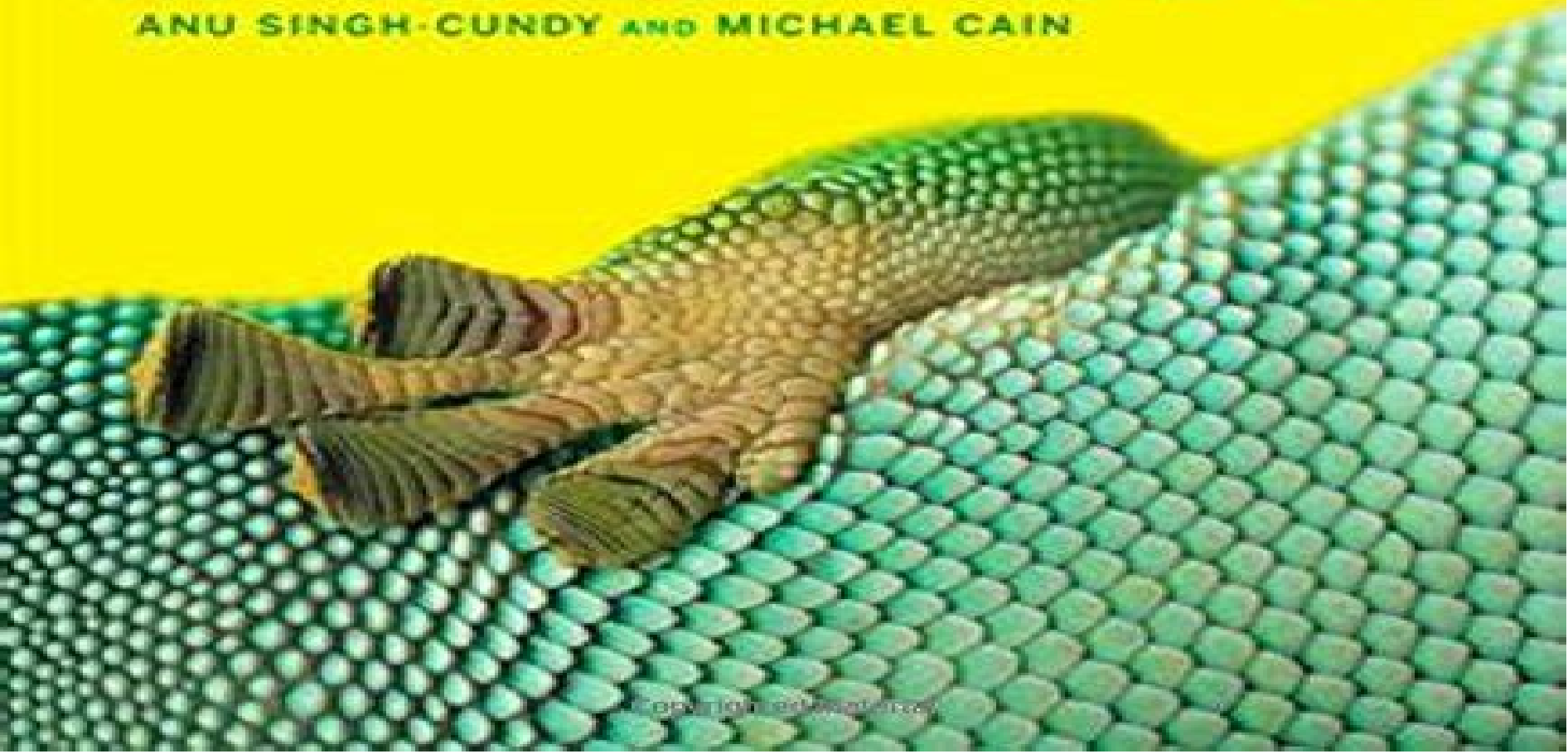


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web step 4 the reactant that produces a smaller amount of product is the limiting reagent mg mg produces less mgo mgo than does o<sub>2</sub> o<sub>2</sub> 3 98 g mgo vs 25 2 g mgo therefore mg is the limiting reagent in this reaction step 5 the reactant that produces a larger amount of product is the excess reagent

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web sep 15 2023 introduction understanding limiting and excess reactants is key in chemistry pogil is an interactive approach that helps students grasp these concepts it provides guided questions and activities here we ll explore pogil answers for limiting and excess reactants and their role in predicting reactions

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web container bodies cylinders tires engines max number of completed cars limiting part a 3 10 9 2 2 engines b 50 12 50 5 c

16 16 16 16 d 4 9 16 6 e 20 36 40

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web please save your changes before editing any questions 5 minutes 1 pt viewing model 3 if 4 moles of hydrogen react with 2 moles of oxygen which of these is the excess reactant is leftover at the end and how many moles of it are leftover hydrogen 2 moles leftover hydrogen 1 mole leftover

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web if you start with 14 82 g of  $\text{Ca(OH)}_2$  and 16 35 g of  $\text{H}_2\text{SO}_4$  a determine the limiting reagent b determine the number of moles of  $\text{H}_2\text{O}$  produced c determine the number of grams of  $\text{CaSO}_4$  produced d determine the number of grams of excess reagent left 1 make sure the equation is balanced

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