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S. No. of Question Paper : 39

Unique Paper Code : 32171101

Name of the Paper : Inorganic Chemistry—I

Name of the Course : B.Sc. (H) Chemistry

Semester : I

Duration : 3 Hours

Maximum Marks : 75

(Write your Roll No. on the top immediately on receipt of this question paper.)

Attempt six questions in all.

Question No. 1 is compulsory.

1. Explain any five of the following with suitable reason : 5×3

- (a) Which is more covalent : NaCl or NaI ?
- (b) Which has the greater bond dissociation energy : O_2 or O_2^+ ?
- (c) All the three N-O bonds in NO_3^- are equal.
- (d) Shape of dz^2 orbital is different from other d -orbitals.
- (e) $BeCl_2$ has zero dipole moment while H_2S has some value.
- (f) Which has greater melting point : *o*-nitrophenol or *p*-nitrophenol ?

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