

Balancing Chemical Equations

Once we have written a chemical equation, we must make sure to balance it! Balancing a chemical equation means to make sure that there are the same number of each type of atom on both sides of the chemical equation. So, if there are five carbons on the left side (reactants), there needs to be five carbons on the right side (products).

So, how do we balance an equation? Follow the following steps...

- Change **only** stoichiometric coefficients – those big numbers in front of the compounds (not subscripts in the formulas – that would change the identity of the compound!)
- Start with most complex
- If found in more than one place on each side, do it last
- Watch out for odd numbers (make them into even ones if needed)
- Check your work!

Now, try to balance the following equations. In addition, name the bolded compounds for some nomenclature practice!

- $\underline{\quad} \text{NH}_4\text{NO}_3(\text{s}) \rightarrow \underline{\quad} \text{N}_2\text{O}(\text{g}) + \underline{2} \text{H}_2\text{O}(\text{g})$
ammonium nitrate & dinitrogen monoxide
- $\underline{\quad} \text{CH}_3\text{CO}_2\text{H}(\text{aq}) + \underline{2} \text{O}_2(\text{g}) \rightarrow \underline{2} \text{CO}_2(\text{g}) + \underline{2} \text{H}_2\text{O}(\text{l})$
oxygen & carbon dioxide
- $\underline{2} \text{Pb}(\text{NO}_3)_2(\text{s}) \rightarrow \underline{2} \text{N}_2\text{O}_4(\text{g}) + \underline{2} \text{PbO}(\text{s}) + \underline{\quad} \text{O}_2(\text{g})$
lead(II) nitrate & dinitrogen tetroxide & lead(II) oxide
- $\underline{\quad} \text{SiO}_2(\text{s}) + \underline{4} \text{HF} \rightarrow \underline{\quad} \text{SiF}_4(\text{g}) + \underline{2} \text{H}_2\text{O}(\text{l})$
silicon dioxide & silicon tetrafluoride
- $\underline{8} \text{H}_2\text{S}(\text{aq}) + \underline{8} \text{Cl}_2(\text{g}) \rightarrow \underline{\quad} \text{S}_8(\text{s}) + \underline{16} \text{HCl}(\text{aq})$
chlorine
- $\underline{\quad} \text{H}_2\text{CO}_3(\text{aq}) + \underline{\quad} \text{Ba}(\text{OH})_2 \rightarrow \underline{\quad} \text{BaCO}_3(\text{s}) + \underline{2} \text{H}_2\text{O}(\text{l})$
barium hydroxide & barium carbonate
- $\underline{\quad} \text{P}_4(\text{s}) + \underline{6} \text{Br}_2(\text{l}) \rightarrow \underline{4} \text{PBr}_3(\text{l})$
phosphorous & phosphorous tribromide
- $\underline{\quad} \text{S}_8(\text{l}) + \underline{4} \text{Cl}_2(\text{g}) \rightarrow \underline{4} \text{S}_2\text{Cl}_2(\text{g})$
disulfur dichloride
- $\underline{2} \text{B}_5\text{H}_9(\text{l}) + \underline{12} \text{O}_2(\text{g}) \rightarrow \underline{5} \text{B}_2\text{O}_3(\text{s}) + \underline{9} \text{H}_2\text{O}(\text{g})$
pentaboron nonahydride & diboron trioxide
- $\underline{3} \text{H}_2\text{SO}_4(\text{aq}) + \underline{2} \text{Al}(\text{OH})_3(\text{s}) \rightarrow \underline{\quad} \text{Al}_2(\text{SO}_4)_3(\text{aq}) + \underline{6} \text{H}_2\text{O}(\text{l})$
aluminum hydroxide & aluminum sulfate
- $\underline{2} \text{Al}(\text{s}) + \underline{3} \text{Br}_2(\text{l}) \rightarrow \underline{2} \text{AlBr}_3(\text{s})$
aluminum bromide
- $\underline{\quad} (\text{NH}_4)_2\text{Cr}_2\text{O}_7(\text{s}) \rightarrow \underline{\quad} \text{N}_2(\text{g}) + \underline{\quad} \text{Cr}_2\text{O}_3(\text{s}) + \underline{4} \text{H}_2\text{O}(\text{l})$
ammonium dichromate & chromium(III) oxide
- $\underline{\quad} \text{S}_8(\text{s}) + \underline{24} \text{F}_2(\text{g}) \rightarrow \underline{8} \text{SF}_6(\text{g})$
sulfur hexafluoride
- $\underline{\quad} \text{Zn}(\text{s}) + \underline{\quad} \text{Pb}(\text{NO}_3)_2(\text{aq}) \rightarrow \underline{\quad} \text{Zn}(\text{NO}_3)_2(\text{aq}) + \underline{\quad} \text{Pb}(\text{s})$
zinc nitrate

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CliffsNotes AP Chemistry Bobrow Test Preparation Services, 2009-02-09. The book itself contains chapter-length subject reviews on every subject tested on the AP Chemistry exam, as well as both sample multiple-choice and free-response questions at each chapter's end. Two full-length practice tests with detailed answer explanations are included in the book.

A New Approach to ICSE Chemistry for Class X (A.Y. 2023-24) Onward V. K. Sally, 2023-05-20. A New Approach to ICSE Chemistry for Class X has been revised in accordance with the latest syllabus prescribed by the Council for the Indian School Certificate Examination, New Delhi, for Class X. The main strength of the book lies in the presentation of scientific

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