

Chapter 9 Test, Form 1

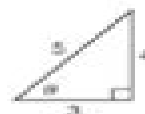
SCORE _____

Write the letter for the correct answer in the blank at the right of each question.

1. Find the value of $\tan \theta$.

A $\frac{4}{3}$
B $\frac{3}{4}$

C $\frac{3}{4}$
D $\frac{4}{3}$



1. A

2. Which equation can be used to find x ?

F $\cos 60^\circ = \frac{4}{x}$
G $\sin 60^\circ = \frac{4}{x}$

H $\tan 60^\circ = \frac{x}{4}$
J $\cot 60^\circ = \frac{4}{x}$



2. G

3. Find the value of x to the nearest tenth.

A 7.8
B 8.0

C 14.4
D 15.0



3. B

4. Find $m\angle P$ to the nearest degree.

F 21°
G 67°

H 23°
J 69°



4. G

5. Rewrite 90° in radian measure.

A $\frac{\pi}{2}$

B $\frac{\pi}{90}$

C $\frac{\pi}{4}$

D $\frac{2}{\pi}$

5. A

6. Rewrite $\frac{\pi}{6}$ radians in degree measure.

F $30\pi^\circ$

G 30°

H 120°

J 60°

6. G

7. Which angle is coterminal with a 90° angle in standard position?

A 540°

B 450°

C -90°

D 270°

7. B

8. Which angle measure is coterminal with the angle $\frac{7\pi}{12}$?

F 15°

G 125°

H 285°

J 465°

8. J

9. Find the arc length for a circle with diameter 10 inches and central angle 60° . Round to the nearest tenth.

A 2.6 in.

B 3.9 in.

C 5.2 in.

D 10.4 in.

9. C

Discovering Geometry Chapter 9 Test Form B

Prof. Sham Tickoo



Discovering Geometry Chapter 9 Test Form B:

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Math Grade 4 Student Edition Randall Inners Charles,1997-06-13 Basic math practice approximately grade 2

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web 33mol o₂ 6mol br₂ 11mol o₂ 18mol br₂ identify the limiting reactant s and excess reactant s the limiting reactant is o₂ since it would yield the least amount of product 18 mol br₂ the excess reactant is c 2 h 3 br 3 since its complete reaction would have yielded up to 54 mol br₂ think about your result

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web step 4 the reactant that produces a smaller amount of product is the limiting reagent mg mg produces less mgo mgo than does o₂ 0.2398 g mgo vs 25.2 g mgo therefore mg is the limiting reagent in this reaction step 5 the reactant that produces a larger amount of product is the excess reagent

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web please save your changes before editing any questions 5 minutes 1 pt viewing model 3 if 4 moles of hydrogen react with 2 moles of oxygen which of these is the excess reactant is leftover at the end and how many moles of it are leftover hydrogen 2 moles leftover hydrogen 1 mole leftover

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web if you start with 14.82 g of ca oh 2 c a o h 2 and 16.35 g of h₂so₄ h 2 s o 4 a determine the limiting reagent b determine the number of moles of h₂o h 2 o produced c determine the number of grams of caso₄ c a s o 4 produced d determine the number of grams of excess reagent left 1 make sure the equation is balanced

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