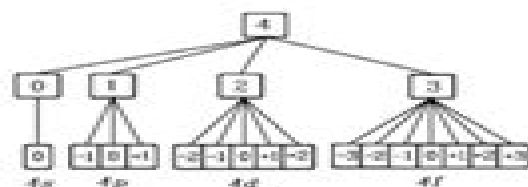


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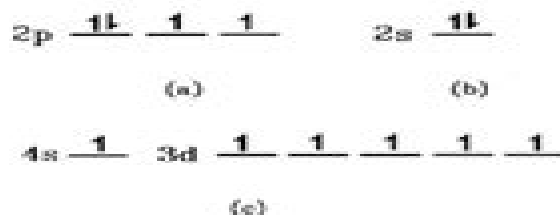
References

- 1.1 (a) Surface where electron probability = 0.
(b) No two electrons in an atom can have the same 4 quantum numbers.
(c) Attraction into a magnetic field by an unpaired electron.

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- 1.5 $5p$.
- 1.7 Size of an orbital.
- 1.9 With parallel spins there is zero probability that the electrons will occupy the same volume of space.
- 1.11 (a) $[\text{Ne}]3s^1$; (b) $[\text{Ar}]4s^2 3d^8$; (c) $[\text{Ar}]4s^1 3d^{10}$.
- 1.13 (a) $[\text{Ar}]$; (b) $[\text{Ar}]$; (c) $[\text{Ar}]3d^9$.
- 1.15 $1+$ and $3+$. Configuration of $[\text{Xe}]6s^2 4f^{14} 5d^1 0p^1$.
- 1.17 $1+$. Configuration of $[\text{Kr}]5s^1 4d^{10}$.
- 1.19



- 1.21 E113: $[Rn]7s^25f^{14}6d^{10}7p^1$
 E113 as +1 ion: $[Rn]7s^25f^{14}6d^{10}$
 E113 as +3 ion: $[Rn]5f^{14}6d^{10}$

Beyond the Basics

- 1.23 9, 5, 121.
- 1.25 There are seven *f*-orbitals. There are two separate ways of depicting them and designating them: the general set and the cubic set.
- 1.27 Hydrogen heads the alkali metal group. Helium heads the alkaline earth metal group.

Chapter 2

Discussion

- 2.1 (a) Lanthanum through lutetium.
(b) Apparent radius of an atom in non-bonded contact with another.
(c) Actual nuclear charge experienced by an electron.
- 2.3 Argon did not fit into any of the then-known groups. Because the table was based on measured atomic mass, argon should have been placed between potassium and calcium.
- 2.5 The long form correctly depicts the order of elements; but the table becomes very elongated.
- 2.7 The *-ium* ending indicates a metal. The ending *-ov* has been used for non-metals.
- 2.9 With nuclei up to 26 protons, nuclear fusion is an exothermic process.
- 2.11 (a) Lead (b) technetium (c) bromine.
- 2.13 Sodium, because it has an odd number of protons.
- 2.15 50.
- 2.17 (a) Several nonmetals have metallic luster.
(b) Diamond has the highest thermal conductivity of all substances.
(c) Graphite is a good electrical conductor in two dimensions.
- 2.19 Potassium. As the effective atomic charge on the outermost electrons increases across a period.
- 2.21 The effective nuclear charge on the *4p* electrons will be increased.

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