#### EXPERIMENT 5 ADSORPTION FROM SOLUTION

#### Introduction

The term adsorption is used to describe the fact that there is a greater concentration of the adsorbed molecules at the surface of the solid than in the bulk solution. In general, one uses solid adsorbents of small size and often with surface imperfections such as cracks and holes which serve to increase the surface area per unit mass greatly over the apparent geometrical area. Such small, porous particles may have specific areas in the range from 10 to 1000 m<sup>2</sup>g<sup>-1</sup>. Some examples of adsorbents commonly used in experiments of this kind are charcoal, silica gel, alumina, zeolites, and molecular sieves. The adsorption from aqueous solutions of acetic cid on charcoal will be investigated in the present experiment.

The type of interaction between the adsorbed molecule and the solid surface varies over a wide range from weak nonpolar van der Waals' forces to strong chemical bonding. Examples of adsorption where ionic or covalent bonding occurs are the adsorption of chloride ions and silver chloride (ionic) or of oxygen gas on metals where oxygen-metal bonds are formed (covalent). In these cases, the process is called chemisorption, and it is generally characterized by high heats of adsorption (from 10 to 100 kcal mol<sup>-1</sup> of gas adsorbed). Chemisorption is highly specific in nature and depends on the chemical properties of both the surface molecules and the adsorbed molecules. Adsorption arising from the weaker van der Waals' and dipole forces is not so specific in character and can take place in any system at low or moderate temperatures. This type of adsorption is called physical adsorption and is usually associated with low heats of adsorption (less than about 10 keal mol'). Physical adsorption forces are similar to those which cause condensation of gases into liquid or solids. When an adsorbing molecule approaches the surface of the solid, there is an interaction between that molecule and the molecule in the surface which tends to concentrate the adsorbing molecules on the surface in much the same way that a gas molecule is condensed onto the surface of bulk liquid. Another respect in which physical adsorption is similar to liquid condensation is the fact that molar heats are of adsorption are of the same order of magnitude as molar heats of vaporization.

The amount adsorbed per gram of solid depends on the specific area of the solid, the equilibrium solute concentration in the solution (or pressure in the case of adsorption from the gas phase), the temperature, and the nature of the molecules involved. From measurements at constant temperature, one can obtain a plot of N, the number of moles adsorbed per gram of solid, versus c, the equilibrium solute concentration. This is called an adsorption isotherm.

Often it is possible to represent experimental results over a limited range by an empirical isotherm suggested by Freundlich:

$$N = Kc^n \tag{1}$$

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