In Table 4 below, [Fe<sup>3+</sup>] after mixing is obtained by using  $M_1V_1 = M_2V_2$  This data is to be transferred to Table 5 below as the initial [Fe<sup>3+</sup>] (column 2).

Similarly in Table 4 of Data sheet [SCN] after mixing obtained using  $M_1V_1 = M_2V_2$ . This data is then to be transferred to Table 5 below as the initial [SCN] (column 3).

Table 4. Initial Concentration Data

Test Tube	[Fe <sup>3+</sup> ] Before Mixing	[Fe <sup>3+</sup> ] After Mixing <sup>1</sup> (Initial [Fe <sup>3+</sup> ])	[SCN <sup>1-</sup> ] Before Mixing	[SCN <sup>1</sup> -] After Mixing <sup>1</sup> (Initial [SCN <sup>1</sup> -])	
F	0.01	0.0025M	0.00113	2.835110-4	
G	0.01	0.0025M	0.00113	5.65 110-4	
н	0.01	0.00514	0.00113	2.825×10-4	
1	0.01	0.005M	0.00113	1.13 × 10 -4	

(Show calculations for initial concentration data on separate sheet)

In Table 5 below, Equilibrium [FeSCN<sup>2+</sup>] is obtained from the Calibration line using absorbance values of solutions F through I.

Equilibrium [Fe3+] = Initial [Fe3+] - Equilibrium [FeSCN2+]

Equilibrium [ SCN ] = Initial [SCN ] - Equilibrium [FeSCN2\*]

Table 5. Equilibrium Constant Calculation

Test Tube	Initial [Fe <sup>3+</sup> ]	Initial [SCN <sup>1-</sup> ]	Equilibrium [FeSCN <sup>2+</sup> ] (from graph)	Equilibrium [Fe <sup>3*</sup> ]	Equilibrium [SCN <sup>1-</sup> ]	K.
F	0.0035	2.825x10-4				
G	0.0025	5.65×10-4		19-		
Н	0.005	2.825x10				
1	0.005	1.13210-4				Same.

(Show calculations for Equilibrium Constant on separate sheet)

Average Equilibrium Constant

# **Determining Equilibrium Constant Lab Report Answers**

**Harold R. Hunt, Toby F. Block** 

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