

Determination of the Solubility-Product Constant for a Sparingly Soluble Salt

A. Preparation of a Calibration Curve

Initial $[\text{CrO}_4^{2-}]$ 0.0024

	Volume of 0.0024 M K_2CrO_4	Total volume	$[\text{CrO}_4^{2-}]$	Absorbance
1.	<u>1 mL</u>	<u>100 mL</u>	<u> </u>	<u>0.122</u>
2.	<u>5 mL</u>	<u>100 mL</u>	<u> </u>	<u>0.625</u>
3.	<u>10 mL</u>	<u>100 mL</u>	<u> </u>	<u>1.288</u>
4.	<u>15 mL</u>	<u>100 mL</u>	<u> </u>	<u>1.955</u>

Molar absorption coefficient for $[\text{CrO}_4^{2-}]$ Error for molar absorption coefficient

B. Determination of the Solubility-Product Constant

	Absorbance	$[\text{CrO}_4^{2-}]$	$[\text{Ag}^+]$	K_{sp} of $\text{Ag}_2\text{CrO}_4 = [\text{CrO}_4^{2-}][\text{Ag}^+]^2$
1.	<u>1.357</u>	<u> </u>	<u> </u>	<u> </u>
2.	<u>1.327</u>	<u> </u>	<u> </u>	<u> </u>
3.	<u>1.549</u>	<u> </u>	<u> </u>	<u> </u>

Determination Of A Solubility Product Lab Answers

Jo Allan Beran



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Electroanalytical and Other Methods Lab Mr. Rohit Manglik, 2024-05-18 EduGorilla Publication is a trusted name in the education sector committed to empowering learners with high quality study materials and resources Specializing in competitive exams and academic support EduGorilla provides comprehensive and well structured content tailored to meet the needs of students across various streams and levels

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