

1. In the list below, circle the properties of acids and box the properties of bases.

Sour taste	Feel slippery	Contains hydroxide ions
Turns litmus blue	Reacts with metals to form H_2	Turns litmus red
Bitter taste	Contains hydrogen ions	Corrosive

2. Complete the chart below:

Acid-Base Definitions		
Type	Acid Definition	Base Definition
Arrhenius	Donates an H^+ ion in solution	Donates an OH^- ion in solution
Bronsted-Lowry	Donates an H^+ ion	Accepts an H^+ ion
Lewis	Accepts an electron pair	Donates an electron pair

3. What term(s) describe a substance that can act as an acid or a base? **Amphoteric or amphiprotic**
4. Neutralization reactions always produce a **SALT** and **WATER**
5. If the $[H^+]$ is greater than the $[OH^-]$, the solution is acidic basic neutral
6. If the $[OH^-]$ is greater than the $[H^+]$, the solution is acidic basic neutral
7. If the $[H^+]$ is equal to the $[OH^-]$, the solution is acidic basic neutral
8. If the pH is 7, the solution is acidic basic neutral
9. If the pH is greater than 7, the solution is acidic basic neutral
10. If the pH is less than 7, the solution is acidic basic neutral
11. How many moles of lithium hydroxide are required to neutralize 19.2 grams of chloric acid?



$$(19.2 \text{ g } HClO_3) \times (1 \text{ mol } HClO_3/84 \text{ g } HClO_3) \times (1 \text{ mol } LiOH/1 \text{ mol } HClO_3) = \underline{0.228 \text{ mol } LiOH}$$

12. How many grams of calcium hydroxide are required to neutralize 132.4 grams of nitric acid?



$$(132.4 \text{ g } HNO_3) \times (1 \text{ mol } HNO_3/63 \text{ g } HNO_3) \times (1 \text{ mol } Ca(OH)_2/2 \text{ mol } HNO_3) \times (74 \text{ g } Ca(OH)_2/1 \text{ mol } Ca(OH)_2) = \underline{77.76 \text{ g } Ca(OH)_2}$$

Acids And Bases Section 3 Answer Key

LL Leslie



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Table of Contents Acids And Bases Section 3 Answer Key

1. Understanding the eBook Acids And Bases Section 3 Answer Key
 - The Rise of Digital Reading Acids And Bases Section 3 Answer Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Acids And Bases Section 3 Answer Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Acids And Bases Section 3 Answer Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from Acids And Bases Section 3 Answer Key
 - Personalized Recommendations
 - Acids And Bases Section 3 Answer Key User Reviews and Ratings
 - Acids And Bases Section 3 Answer Key and Bestseller Lists
5. Accessing Acids And Bases Section 3 Answer Key Free and Paid eBooks
 - Acids And Bases Section 3 Answer Key Public Domain eBooks
 - Acids And Bases Section 3 Answer Key eBook Subscription Services
 - Acids And Bases Section 3 Answer Key Budget-Friendly Options
6. Navigating Acids And Bases Section 3 Answer Key eBook Formats

- ePub, PDF, MOBI, and More
 - Acids And Bases Section 3 Answer Key Compatibility with Devices
 - Acids And Bases Section 3 Answer Key Enhanced eBook Features
7. Enhancing Your Reading Experience
- Adjustable Fonts and Text Sizes of Acids And Bases Section 3 Answer Key
 - Highlighting and Note-Taking Acids And Bases Section 3 Answer Key
 - Interactive Elements Acids And Bases Section 3 Answer Key
8. Staying Engaged with Acids And Bases Section 3 Answer Key
- Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Acids And Bases Section 3 Answer Key
9. Balancing eBooks and Physical Books Acids And Bases Section 3 Answer Key
- Benefits of a Digital Library
 - Creating a Diverse Reading Collection Acids And Bases Section 3 Answer Key
10. Overcoming Reading Challenges
- Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Acids And Bases Section 3 Answer Key
- Setting Reading Goals Acids And Bases Section 3 Answer Key
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Acids And Bases Section 3 Answer Key
- Fact-Checking eBook Content of Acids And Bases Section 3 Answer Key
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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