

Acid/Base Titration Lab

Background: Titration involves the addition of a solution whose concentration is known to a solution whose concentration is unknown. The volume of the known solution required to react completely with a known volume of the solution whose concentration is being determined is measured. An indicator is added to the solution to mark the point at which the two quantities reach equivalence.

In this experiment sodium hydroxide of known concentration will be added to a known volume of hydrochloric acid to determine the molarity of the sodium hydroxide. Phenolphthalein will be used as the indicator. Phenolphthalein is pink in a basic solution and clear in an acidic solution.

Objective: To determine the molarity of an unknown concentration of hydrochloric acid.

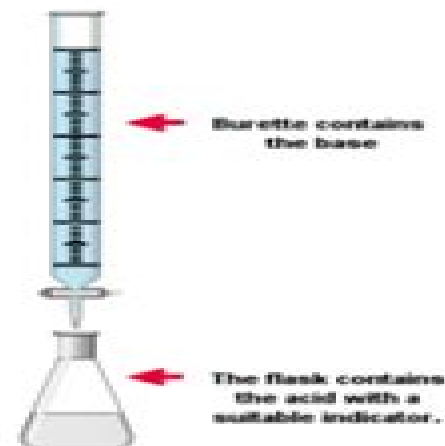
Safety: Wear goggles and aprons – both acids and bases are corrosive and can cause eye and/or skin damage. *Do not pour the NaOH above eye level. Lower the buret below eye level before filling it.* Water soluble solutions can be poured down the sink.

Materials:

Ring stand	1 Erlenmeyer flask	phenolphthalein	2 beakers	Buret	DI water	0.5 M NaOH
Funnel	Clamp	10 mL syringe	unknown acid	100 mL graduated cylinder		

Procedure:

1. Set up the ring stand, buret, and buret clamp. Make sure that the buret remains vertical at all times. See image to the right.
2. Add some DI water to the buret to check for leaks. If there are no leaks, drain the DI water completely from the buret.
3. Carefully add the NaOH to the buret. You will need to drain some of the NaOH from the buret to fill the tip. Fill the buret so that the meniscus of the NaOH is sitting on the 0.00 mL mark with the tip filled.
4. Using the graduated cylinder add about 60 mL of DI water to a 250 mL Erlenmeyer flask (or 20 mL to a 125 mL Erlenmeyer flask).
5. Using the syringe add exactly 10 mL of the acid unknown to the water in the Erlenmeyer flask. Make sure that you write down the letter of your unknown.
6. Add 2 – 3 drops of phenolphthalein to the acid solution in the Erlenmeyer flask.
7. Place the Erlenmeyer under the buret. Place a sheet of white paper under the Erlenmeyer.
8. You are now ready to begin titrating. Carefully add about 1 mL of the NaOH from the buret to the Erlenmeyer flask and swirl to mix. Continue adding 1 mL increments until you begin to see a pink color. The pink color should appear briefly and then quickly disappear.
9. Once you have seen the pink color, begin adding the NaOH slowly and in smaller amounts. Remember to swirl the flask after each addition of NaOH. As you near the endpoint, the pink color should stay longer, but will still disappear when the flask is swirled. The longer the pink color stays, the smaller your next



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